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Tokyo, Japan, solemnly and sincerely declare:

That I have a thorough knowledge of Japanese  
and English languages; and

That the attached pages contain a correct  
translation into English of the specification of the  
following Japanese Application:

APPLICATION  
NUMBERDATE OF  
APPLICATION

284599/2001(Pat.) 19/SEP/2001

Applicant(s)

CANON KABUSHIKI KAISHA

Signed this 19th day of September, 2003.  
Ryuichi YAMADA

**PATENT OFFICE**  
**JAPANESE GOVERNMENT**

**This is to certify that the annexed is a true  
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<b><u>APPLICATION NUMBER</u></b>	<b><u>DATE OF APPLICATION</u></b>
<b>284599/2001(Pat.)</b>	<b>19/SEP/2001</b>

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**19/OCT/2001**

**Director-General,  
Patent Office : KOZO OIKAWA (Seal)**

**Certificate No. 3092040/2001**

<FILE NO.> 4557001

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<DOCUMENT> PATENT APPLICATION

<FILE NO.> 4557001

<FILING DATE> 19/SEP/2001

<DIRECTED TO> The Director General of the Patent Office

<INTERNATIONAL CLASSIFICATION> H05B 33/00

<TITLE OF THE INVENTION>

LUMINESCENCE DEVICE, DISPLAY APPARATUS AND METAL  
COORDINATION COMPOUND FOR THE LUMINESCENCE DEVICE

<NUMBER OF CLAIMS> 26

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<CLAIM FOR PRIORITY>	
<APPLICATION NO.>	292492/2000
<FILING DATE>	26/SEP/2000
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<FILING DATE>	26/SEP/2000
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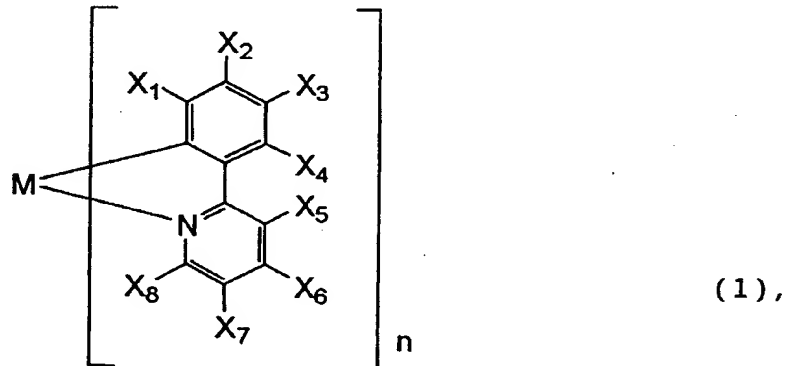
Specification

[Title of the Invention]

Luminescence Device, Display Apparatus and  
Metal Coordination Compound for the  
Luminescence Device

[Claims]

1. A luminescence device, comprising: an organic compound layer comprising a metal coordination compound represented by the following formula (1):



{wherein M denotes Ir, Rh or Pd; n is 2 or 3; and X1 to X8 independently denote hydrogen atom or a substituent [selected from the group consisting of halogen atom; nitro group; trifluoromethyl group trialkylsilyl group (which alkyl groups are independently a linear or branched alkyl group each having 1 - 8 carbon atoms); and a linear or branched

alkyl group having 2 - 20 carbon atoms wherein one or at least two non-neighboring methylene groups which can be replaced with -O-, -S-, -CO-, -CO-O-, -O-CO-, -CH=CH- or -C≡C- and hydrogen atom in said alkyl group can be replaced with fluorine atom)]; with the proviso that at least one of X1 to X8 is a substituent other than hydrogen atom, and X2 and X3 cannot be fluorine atom at the same time}.

2. A device according to Claim 1, wherein in the formula (1), at least two of X1 to X8 are substituents other than hydrogen atom.

3. A device according to Claim 1 or 2, wherein in the formula (1), at least one of X5 to X8 is a substituent other than hydrogen atom.

4. A device according to any one of Claims 1 - 3, wherein in the formula (1), at least two of X1 to X4 are substituents other than hydrogen atom.

5. A device according to any one of Claims 1 - 4, wherein in the formula (1), at least one of X2, X3 and X4 has a Hammett's substituent constant of at least 0.2, and the metal coordination compound exhibits a peak emission wavelength in toluene at a normal temperature of 25 °C, of at most 490 nm.

6. A device according to any one of Claims 1 - 5, wherein in the formula (1), X2, X3 and X4 provides a sum of Hammett's substituent constant of at least 0.41, and the metal coordination compound exhibits a peak emission wavelength in toluene at a normal temperature of 25 °C, of at most 490 nm.

7. A device according to Claim 6, wherein the sum of Hammett's substituent constant is at least 0.50.

8. A luminescence device, comprising: an organic compound layer comprising a metal coordination compound which has a substituent having a Hammett's substituent constant of at least 0.2, and exhibits a peak emission wavelength, in toluene at a normal temperature of 25 °C, of at most 490 nm.

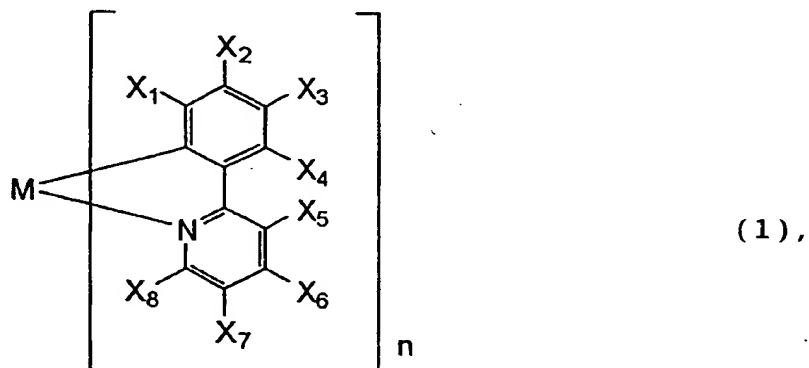
9. A device according to Claim 8, wherein said metal coordination compound has a plurality of substituents each having a Hammett's substituent constant of at least 0.2.

10. A device according to any one of Claims 1 - 9, further comprising two electrodes oppositely disposed to sandwich the organic compound layer to

constitute an electroluminescence device, wherein a voltage is applied between the electrodes to produce luminescence.

11. A display apparatus, comprising: a luminescence device according to any one of Claims 1 - 9 and a portion for driving the luminescence device.

12. A metal coordination compound for a luminescence device, represented by the following formula (1):



{wherein M denotes Ir, Rh or Pd; n is 2 or 3; and X1 to X8 independently denote hydrogen atom or a substituent [selected from the group consisting of halogen atom; nitro group; trifluoromethyl group trialkylsilyl group (which alkyl groups are independently a linear or branched alkyl group each having 1 - 8 carbon atoms; and a linear or branched alkyl group having 2 - 20 carbon atoms wherein one or at least two non-



neighboring methylene groups which can be replaced with -O-, -S-, -CO-, -CO-O-, -O-CO-, -CH=CH- or -C≡C- and hydrogen atom in said alkyl group can be replaced with fluorine atom)]; with the proviso that at least one of X1 to X8 is a substituent other than hydrogen atom, and X2 and X3 cannot be fluorine atom at the same time}.

13. A compound according to Claim 12, wherein in the formula (1), at least two of X1 to X8 are substituents other than hydrogen atom.

14. A compound according to Claim 12 or 13, wherein in the formula (1), at least one of X5 to X8 is a substituent other than hydrogen atom.

15. A compound according to any one of Claims 12 - 14, wherein in the formula (1), at least two of X1 to X4 are substituents other than hydrogen atom.

16. A compound according to any one of Claims 12 - 15, wherein in the formula (1), at least one of X2, X3 and X4 has a Hammett's substituent constant of at least 0.2, and the metal coordination compound exhibits a peak emission wavelength, in toluene at a normal temperature of 25 °C, of at most 490 nm.

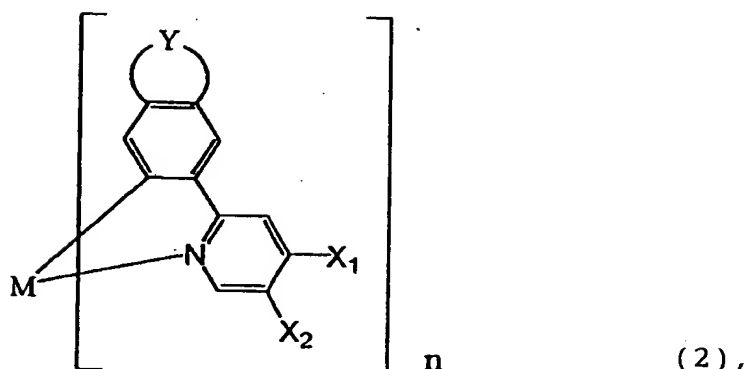
17. A compound according to any one of Claims 12 - 16, wherein in the formula (1), X2, X3 and X4 provides a sum of Hammett's substituent constant of at least 0.41, and the metal coordination compound exhibits a peak emission wavelength, in toluene at a normal temperature of 25 °C, of at most 490 nm.

18. A compound according to Claim 17, wherein the sum of Hammett's substituent constant is at least 0.50.

19. A metal coordination compound, which has a substituent having a Hammett's substituent constant of at least 0.2, and exhibits a peak emission wavelength, in toluene at a normal temperature of 25 °C, of at most 490 nm.

20. A compound according to Claim 19, which has a plurality of substituents each having a Hammett's substituent constant of at least 0.2.

21. A luminescence device, comprising: an organic compound layer comprising a metal coordination compound represented by the following formula (2):



[wherein M denotes Ir, Pt, Rh or Pd; n is 2 or 3; Y denotes an alkylene group having 2 - 4 carbon atoms (wherein one or at least two non-neighboring methylene groups may be replaced with -O-, -S- or -CO- and hydrogen atom in said alkylene group may be replaced with a linear or branched alkyl group having 1 - 10 carbon atoms); and X1 and X2 independently denote hydrogen atom; halogen atom; nitro group; trialkylsilyl group having 1 - 8 carbon atoms; or a linear or branched alkyl group having 1 - 20 carbon atoms (wherein one or at least two non-neighboring methylene groups may be replaced with -O-, -S-, -CO-, -CO-O-, -O-CO-, -CH=CH- or -C≡C- and capable of including hydrogen atom which can be replaced with fluorine atom)].

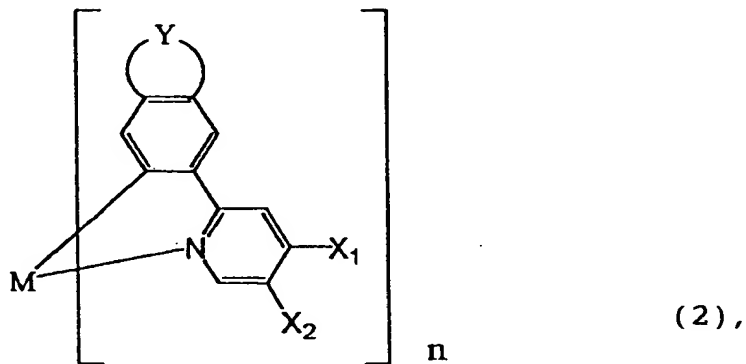
22. A device according to Claim 21, wherein in the formula (2), at least one of X1 to X2 is hydrogen atom.

23. A device according to Claim 21 or 22, further

comprising two electrodes oppositely disposed to sandwich the organic compound layer to constitute an electroluminescence device, wherein a voltage is applied between the electrodes to produce luminescence.

24. A display apparatus, comprising: a luminescence device according to any one of Claims 21 - 23 and a portion for driving the luminescence device.

25. A metal coordination compound for a luminescence device, represented by the following formula (2):



[wherein M denotes Ir, Pt, Rh or Pd; n is 2 or 3; Y denotes an alkylene group having 2 - 4 carbon atoms (wherein one or at least two non-neighboring methylene groups may be replaced with -O-, -S- or -CO- and hydrogen atom in said alkylene group may be replaced with a linear or branched alkyl group having 1 - 10 carbon atoms); and X1 and X2 independently denote hydrogen

atom; halogen atom; nitro group; trialkylsilyl group having 1 - 8 carbon atoms; or a linear or branched alkyl group having 1 - 20 carbon atoms (wherein one or at least two non-neighboring methylene groups may be replaced with -O-, -S-, -CO-, -CO-O-, -O-CO-, -CH=CH- or -C≡C- and capable of including hydrogen atom which can be replaced with fluorine atom)].

26. A compound according to Claim 25, wherein in the formula (2), at least one of X1 to X2 is hydrogen atom.

[Detailed Description of the Invention]

[0001]

[Technical Field to which the Invention Pertains]

The present invention relates to a metal coordination compound for a luminescence device, a luminescence device using the compound, and a display apparatus. More specifically, the present invention relates to an organic electroluminescence device employing a metal coordination compound represented by the above-mentioned formula (1) or (2) as a luminescence material, and a display apparatus.

[0002]

[Prior Art]

An organic EL (electroluminescence) device has been extensively studied as a luminescence device

with a high responsiveness and high efficiency. The organic EL device generally has a basic structure as shown in Figure 1(a), (b) (e.g., as described in "Macromol. Symp.", 125, pp. 1 - 48 (1997)).

[0003]

As shown in Figure 1, the EL device is generally constituted by disposing, on a transparent substrate 15, a plurality of organic film layers between a transparent electrode 14 and a metal electrode 11.

In Figure 1(a), the organic layers includes a luminescence layer 12 and a hole transport layer 13.

As transparent electrode 14, a film of ITO (indium tin oxide) having a larger work function is used so as to ensure a good hole injection performance from the transparent electrode 14 into the hole transport layer 13. As the metal electrode 11, a layer of metal material such as aluminum, magnesium, alloys thereof, etc., having a smaller work function is used so as to ensure a good electron injection performance into the organic layer(s).

These electrodes may be formed in a thickness of 50 - 200 nm.

[0005]

The luminescence layer 12 may be formed of, e.g., aluminum quinolinol complex (representative example thereof may include Alq3 described

hereinafter) having an electron transporting characteristic and a luminescent characteristic. The hole transport layer 13 may be formed of, e.g., triphenyldiamine derivative (representative example thereof may include  $\alpha$ -NPD described hereinafter) having an electron donating characteristic.

[0006]

The above-described EL device exhibits an electrical rectification characteristic, so that when an electric field is applied between the metal electrode 11 as a cathode and the transparent electrode 14 as an anode, electrons are injected from the metal electrode 11 into the luminescence layer 12 and holes are injected from the transparent electrodes 15.

[0007]

The thus-injected holes and electrons are recombined within the luminescence layer 12 to produce excitons, thus producing luminescence. At that time, the hole transport layer 13 functions as an electron-blocking layer to increase a recombination efficiency at the boundary between the luminescence layer 12 and the hole transport layer 13, thus enhancing a luminescence efficiency.

[0008]

Further, in Figure 1(b), an electron transport layer 16 is disposed between the metal

electrode 11 and the luminescence layer 12 shown in Figure 1(a), whereby an effective carrier blocking performance can be ensured by separating functions of luminescence, electron transport and hole transport, thus allowing efficient luminescence. The electron transport layer 16 may be formed of, e.g., oxadiazole derivatives.

[0009]

Heretofore, in ordinary organic EL devices, fluorescence produced during a transition of luminescent center molecule from a singlet excited state to a ground state has been used as luminescence. On the other hand, not the above fluorescence (luminescence) via singlet exciton, phosphorescence (luminescence) via triplet exciton has been studied for use in organic EL device as described in, e.g., Document 1: "Improved energy transfer in electrophosphorescent device" (D.F. O'Brien et al., Applied Physics Letters, Vol. 74, No. 3, pp. 442 - (1999)) and Document 2: "Very high-efficiency green organic light-emitting devices based on electrophosphorescence" (M.A. Baldo et al., Applied Physics Letters, Vol. 75, No. 1, pp. 4 - 1999)).

[0010]

In these documents, a four layer-structure of organic layers as shown in Figure 1(c) is principally used. The structure includes a hole transfer layer



13, a luminescence layer 12, an exciton diffusion-prevention layer 17, and an electron transport layer 16 successively formed in this order from the anode side. The materials used therefor are carrier transporting materials and phosphorescent materials shown below. Abbreviations for the respective materials are follows.

Alq3: aluminum-quinolinol complex,

$\alpha$ -NPD: N4,N4'-di-naphthalene-1-yl-N4,N4'-diphenyl-biphenyl-4,4'-diamine,

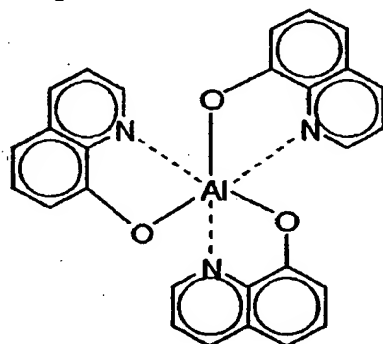
CBP: 4,4'-N,N'-dicarbazole-biphenyl,

BCP: 2,9-dimethyl-4,7-diphenyl-1,10-phenanthroline,

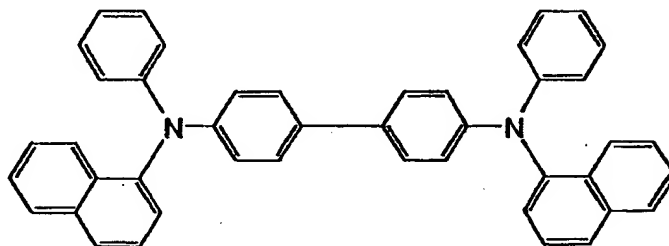
PtEOP: platinum-octaethyl porphine complex, and

Ir(ppy)<sub>3</sub>: iridium-phenylpyridine complex.

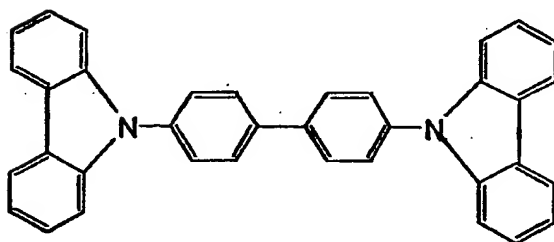
[0011]



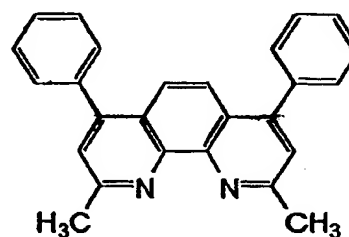
Alq3



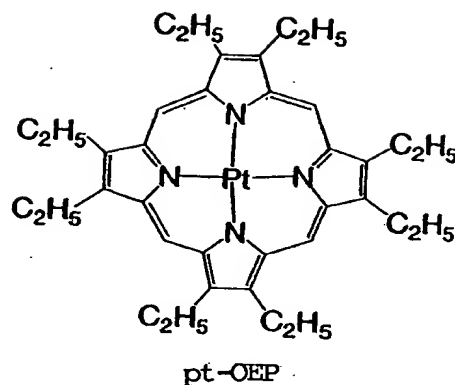
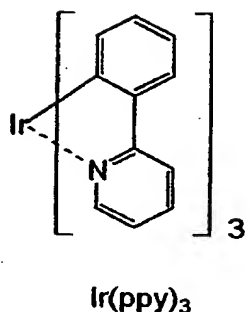
$\alpha$ -NPD



CBP



BCP



[0012]

In the above Documents 1 and 2, higher efficiencies have been achieved by using a hole transport layer 13 of  $\alpha$ -NPD, an electron transport layer 16 of Alq3, an exciton diffusion-prevention layer 17 of BPC, and a luminescence layer 12 of a mixture of CPB as a host material with  $\text{Ir(ppy)}_3$  or PtOEP doped into CBP at a concentration of ca. 6 wt. %.

[0013]

The reason why the phosphorescence (luminescence) material has particularly attracted notice is that the phosphorescence material is expected to provide a higher luminescence efficiency in principle. The reason is as follows. Excitons produced by recombination of carriers comprise singlet excitons and triplet excitons presented in a ratio of 1:3. Fluorescence produced during the transition from

the singlet excited state to the ground state is utilized as luminescence in the conventional organic EL devices. However, a resultant luminescence efficiency is 25 % (being upper limit) based on all the produced excitons in principle.

However, if phosphorescence produced during transition from the triplet excited state is employed, a resultant luminescence efficiency is expected to be at least three times (that of the case of fluorescence) in principle. In addition thereto, if intersystem crossing from the singlet excited state which is a higher energy level to the triplet excited state is taken into consideration, the luminescence efficiency of phosphorescence can be expected to be 100 %, which is four times that of fluorescence, in principle.

[0014]

The use of phosphorescence based on transition from the triplet excited state has also been proposed in, e.g., Japanese Laid-Open Patent Application (JP-A) 11-329739 (organic EL device and production process thereof), JP-A 11-256148 (luminescent material and organic EL device using the same) and JP-A 8-319482 (organic electroluminescent device).

[0015]

An unsubstituted iridium-phenylpyrimidine

complex and six species of substituted iridium-phenylpyrimidine complexes have been described in "Inorg. Chem.", 1991, 30, 1685 - 1687. However, in the document, description as to a peak wavelength of emission spectrum in solution on the shortest wavelength side is merely formed and there is no description with respect to the organic EL device.

[0016]

An iridium-phenylpyrimidine complex, closer to the present invention, having a methyl substituent has been described in "Preprint for the 61-th Scientific Lecture of the Applied Physics Society of Japan", the third volume, P.1117, 6p-ZH-1 (2000).

[0017]

Further, an iridium-phenylpyrimidine complex, closer to the present invention, having 4-, 5-fluorine substituents (herein, referred to as a "metal coordination compound A" has been described in "Polymer Preprints", 41(1), pp. 770 - 771 (2000).

[0018]

[Problems to be Solved by the Invention]

The above-mentioned organic EL devices utilizing phosphorescence have accompanied with a problem of luminescent deterioration particularly in an energized state. The reason for luminescent deterioration has not been clarified as yet but may be attributable to such a phenomenon that the life of

triplet exciton is generally longer than that of singlet exciton by at least three digits, so that molecule is placed in a higher-energy state for a long period to cause reaction with ambient substance, formation of exciplex or excimer, change in minute molecular structure, structural change of ambient substance, etc.

[0019]

Anyway, the phosphorescence luminescence device is expected to provide a higher luminescence efficiency as described above, while the device is accompanied with the problem of luminescent deterioration in energized state.

[0020]

Accordingly, an object of the present invention is to provide a luminescence device and a display apparatus, capable of providing a high-efficiency luminescence at a high brightness (or luminance) for a long period (less deterioration in luminescence in energized state). Another object of the present invention is to provide a metal coordination compound used for such luminescence device and display apparatus.

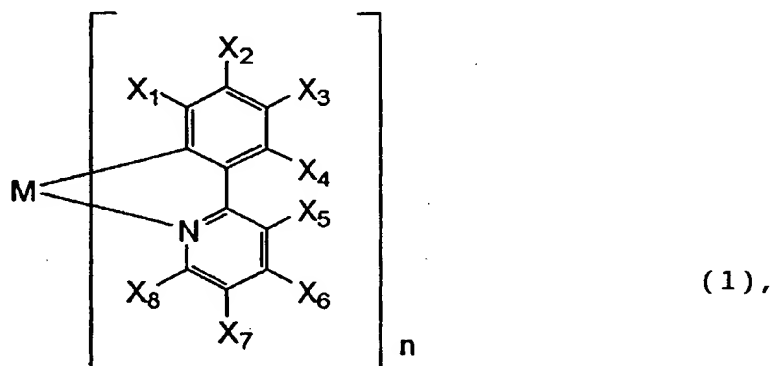
Incidentally, although values are different depending on required performances, for example, a luminescence efficiency of 1 cd/W or above in response to an inputted current value may be referred to as a

higher efficiency, and a luminance half-life time of 500 hours or above in the case of luminescence at an initial luminance of  $100 \text{ cd/m}^2$  may be referred to as a longer high-luminance period of time or less deterioration in energized state.

[0021]

[Means for Solving the Problems]

More specifically, the present invention provided a luminescence device, comprising: an organic compound layer comprising a metal coordination compound represented by the following formula (1):



[0023]

{wherein M denotes Ir, Rh or Pd; n is 2 or 3; and X1 to X8 independently denote hydrogen atom or a substituent [selected from the group consisting of halogen atom; nitro group; trifluoromethyl group trialkylsilyl group (which alkyl groups are independently a linear or branched alkyl group each having 1 - 8 carbon atoms); and a linear or branched

alkyl group having 2 - 20 carbon atoms wherein one or at least two non-neighboring methylene groups which can be replaced with -O-, -S-, -CO-, -CO-O-, -O-CO-, -CH=CH- or -C≡C- and hydrogen atom in said alkyl group can be replaced with fluorine atom)]; with the proviso that at least one of X1 to X8 is a substituent other than hydrogen atom, and X2 and X3 cannot be fluorine atom at the same time}.

[0024]

In the luminescence device of the present invention, in the formula (1), at least two of X1 to X8 may preferably be substituents other than hydrogen atom; at least one of X5 to X8 may preferably be a substituent other than hydrogen atom; and at least two of X1 to X4 may preferably be substituents other than hydrogen atom.

[0025]

Further, in the formula (1), it is preferable that at least one of X2, X3 and X4 has a Hammett's substituent constant of at least 0.2, and the metal coordination compound exhibits a peak emission wavelength in toluene at a normal temperature of 25 °C, of at most 490 nm. In the formula (1), X2, X3 and X4 may preferably provide a sum of Hammett's substituent constant of at least 0.41, more preferably at least 0.50, and the metal coordination compound may preferably exhibit a peak emission wavelength in

toluene at a normal temperature of 25 °C, of at most 490 nm.

[0026]

The present invention provides a luminescence device, comprising: an organic compound layer comprising a metal coordination compound which has a substituent having a Hammett's substituent constant of at least 0.2, and exhibits a peak emission wavelength, in toluene at a normal temperature of 25 °C, of at most 490 nm. The metal coordination compound may preferably have a plurality of substituents each having a Hammett's substituent constant of at least 0.2.

[0027]

Further, the aforementioned luminescence device may preferably further comprise two electrodes oppositely disposed to sandwich the organic compound layer to constitute an electroluminescence device, wherein a voltage is applied between the electrodes to produce luminescence.

[0028]

Further, the present invention provide a display apparatus, comprising: the above-mentioned luminescence device and a portion for driving the luminescence device.

[0029]

Further, the present invention provides a

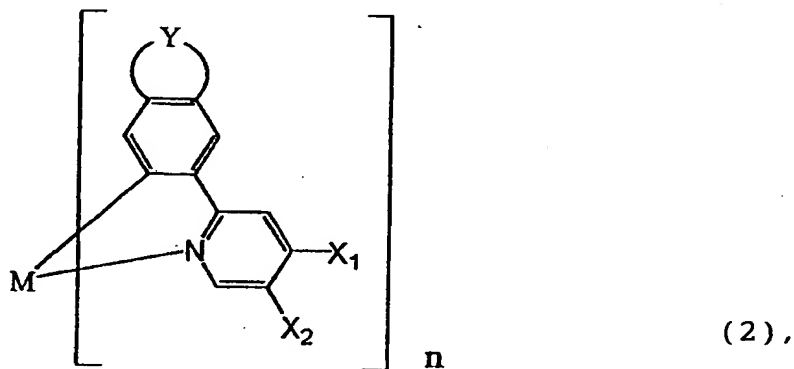


metal coordination compound represented by the aforementioned formula (1) for a luminescence device, and a metal coordination compound, which has a substituent having a Hammett's substituent constant of at least 0.2, and exhibits a peak emission wavelength, in toluene at a normal temperature of 25 °C, of at most 490 nm.

[0030]

Further, the present invention provides another luminescence device, comprising: an organic compound layer comprising a metal coordination compound represented by the following formula (2):

[0031]



[0032]

[wherein M denotes Ir, Pt, Rh or Pd; n is 2 or 3; Y denotes an alkylene group having 2 - 4 carbon atoms (wherein one or at least two non-neighboring methylene groups may be replaced with -O-, -S- or -CO- and hydrogen atom in said alkylene group may be replaced with a linear or branched alkyl group having 1 - 10

carbon atoms); and X1 and X2 independently denote hydrogen atom; halogen atom; nitro group; trialkylsilyl group having 1 - 8 carbon atoms; or a linear or branched alkyl group having 1 - 20 carbon atoms (wherein one or at least two non-neighboring methylene groups may be replaced with -O-, -S-, -CO-, -CO-O-, -O-CO-, -CH=CH- or -C≡C- and capable of including hydrogen atom which can be replaced with fluorine atom)].

[0033]

In another luminescence device according to the present invention, in the formula (2), at least one of X1 to X2 may preferably be hydrogen atom.

[0034]

Further, another luminescence device may preferably further comprise two electrodes oppositely disposed to sandwich the organic compound layer to constitute an electroluminescence device, wherein a voltage is applied between the electrodes to produce luminescence.

[0035]

Further, the present invention provides a display apparatus, comprising: another luminescence device described above and a portion for driving the luminescence device.

[0036]

Further, the present invention provides a

metal coordination compound for a luminescence device, represented by the aforementioned formula (2).

[0037]

[Embodiments for Practicing the Invention]

In the case where a luminescence layer is formed of a carrier transporting host material and a phosphorescent guest material, a process of emission of light from the triplet exciton may generally involve the following steps:

- (1) transport of electron and hole within a luminescence layer,
- (2) formation of exciton of the host material,
- (3) transmission of excited energy between host material molecules,
- (4) transmission of excited energy from the host material molecule to the guest material molecule,
- (5) formation of triplet exciton of the guest material, and
- (6) phosphorescence produced during transition from the triplet excited state to the ground state of the guest material.

[0038]

In the respective steps, desired energy transmission and luminescence may be caused based on various deactivation and competition.

[0039]

In order to improve a luminescence efficiency

of the EL device, a luminescence center material per se is required to provide a higher yield of luminescence quantum. However, an efficient energy transfer between host material molecules and/or between host material molecule and guest material molecule is also an important factor. Further, the case of the luminescent deterioration in energized state has not been clarified as yet but may presumably relate to the luminescent center material per se or an environmental change thereof by its ambient molecular structure.

[0040]

For this reason, the present inventors have extensively studied on various metal coordination compounds and as a result, have found that the organic electroluminescence device using the metal coordination compound represented by the above-mentioned formula (1) (hereinafter, sometimes referred to as a "material (1)" as the luminescence center material allows a high-efficiency luminescence with a high brightness (luminance) for a long period and less luminescent deterioration in energized state.

[0041]

The metal coordination compound of formula (1) may preferably have substituents X1 to X8 in which at least two of X1 to X8 are substituents other than hydrogen atom. Further, in the formula (1), at least

one of X5 to X8 may preferably be a substituent other than hydrogen atom and/or at least two of X1 to X4 may preferably be substituents other than hydrogen atom.

[0042]

The metal coordination compound being the material (1) has a phosphorescent property and is assumed to have a lowest excited state comprising, as a triplet excited state, a metal-to-ligand charge transfer (MLCT\*) state or  $\pi$ - $\pi^*$  excited state. The phosphorescent emission of light (phosphorescence) is produced during the transition from these states to the ground state.

[0043]

The material (1) provides a higher phosphorescence yield of 0.1 - 0.9 and a shorter phosphorescence life of 1 - 60  $\mu$ sec.

More specifically, the shorter phosphorescence life is a condition for providing a resultant EL device with a higher luminescence efficiency. This is because the longer phosphorescence life increases molecules placed in their triplet excited state which is a waiting state for phosphorescence, thus resulting in a problem of lowering the resultant luminescence efficiency particularly at a higher current density. The material (1) is a suitable luminescent material for an EL device with a higher phosphorescence yield and a

shorter phosphorescence life.

Incidentally, a phosphorescence yield ( $F(m)$ ) is obtained based on the following equation:

$$F(m)/F(s) = (S(m)/S(s)) \times (A(s)/A(m)),$$
 wherein  $F$  represents a phosphorescence yield,  $S$  represents an integrated intensity of (photo-)excited emission spectrum of the luminescent material,  $A$  represents an absorption spectrum of an excited light wavelength, and  $S$  represents a material with a known phosphorescence yield, and  $m$  represents an unknown material.

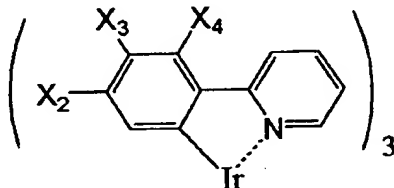
[0044]

Further, it has been found that it is possible to control an emission wavelength of the metal coordination compound as the material (1) by appropriately modifying the substituents  $X_1$  to  $X_8$  thereof. In this regard, as a result of our investigation on various phosphorescence metal coordination compounds in order to search a blue luminescence material required to have a peak (maximum) emission wavelength of at most 490 nm, we have found that it is very effective to introduce a substituent having a Hammett's substituent constant of at least 0.2 into the metal coordination compound of formula (1) in order to provide a shorter peak emission wavelength.

[0045]

Hereinbelow, explanation will be made by taking the metal coordination compound represented by the formula (1) of the present invention as an example. We focus attention on a carbon atom connected to iridium metal of a metal coordination compound represented by the formula shown below, and have investigated a relationship between Hammett's substituent constants ( $\sigma$ ) of substituents X2, X3 and X4, and peak emission wavelengths ( $\lambda_{PE}$ ) of the metal coordination compound in toluene at room temperature, i.e., at 25 °C.

[0046]



[0047]

A Hammett's substituent constant  $\sigma_m$  for meta-position was used for the substituents X2 and X4 and a Hammett's substituent constant  $\sigma_p$  for para-position was used for the substituent X3. When a plurality of substituents were present, a sum of the respective substituent constants ( $\sigma_m$  and  $\sigma_p$ ) was used (as a Hammett's substituent constant  $\sigma$ ). Here, Hammett's substituent constants  $\sigma_m$  and  $\sigma_p$  described on pages 96 - 103 (Table 1) of "Correlation between Structure and

Activation of Drugs", Chemical Region Extra Edition 122, issued by Nanko-do (Japan) were used. A part of the substituent constants described therein is shown in Table 1 below.

[0048]

[Table 1]

Substituent	Hammett's $\sigma_p$	Constant $\sigma_m$
F	0.06	0.34
Cl	0.23	0.37
CF <sub>3</sub>	0.54	0.43

[0049]

For example, when the Hammett's substituent constant  $\sigma$  of Example Compound No. (121) ( $X_2 = F$ ,  $X_3 = CF_3$ ) is calculated,  $0.34 + 0.54 = 0.88$ . In a similar manner, Hammett's substituent constants  $\sigma$  of Ex. Comp. Nos. (1), (32), (122) and (111) and the metal coordination compound A are calculated. The results are shown in Table 2 below together with corresponding peak (maximum) emission wavelength  $\lambda_{PE}$  in toluene. The results of Table 2 are plotted by taking the Hammett's substituent constant  $\sigma$  on the axis of abscissa and the peak emission wavelength  $\lambda_{PE}$  on the axis of ordinate (Figure 2).

[0050]



[Table 2]

Compound	$\sigma$	PE (nm)
Ex. Comp. No. (1)	0.06	522
Metal coordination compound A	0.40	505
Ex. Comp. No. (32)	0.54	487
Ex. Comp. No. (122)	0.68	471
Ex. Comp. No. (121)	0.88	466
Ex. Comp. No. (111)	0.91	479

[0051]

As apparent from Figure 2, introduction of substituent(s) having a larger Hammett's substituent constant is very effective to shorten the peak emission wavelength. Further, the metal coordination compound having the sum of peak emission wavelengths of at least 0.41, particularly at least 0.50 is suitable as the blue luminescent material. A similar effect can be expected also for metal coordination compounds other than the metal coordination compound of the present invention. From also the viewpoints as described above, the metal coordination compound of the present invention is a suitable luminescent material for the EL device.

[0052]

Further, as shown in Examples appearing

hereinafter, it has been clarified that the (metal coordination) compound of the present invention also has an excellent stability in a continuous energization test.

This may be attributable to introduction of particular substituents as a feature of the present invention allowing control of intermolecular interaction with a host luminescent material and suppression of formation of associated exciton leading to thermal inactivation, thus minimizing quenching to improve device characteristics. Further, the methyl group of methyl-substituted iridium-phenylpyrimidine complex described in the above-mentioned "Preprint for the 61-th Scientific Lecture of the Applied Physics Society of Japan", the third volume, P. 1117, 6p-ZH-1 (2000), has a smaller bulkiness than ethyl group and methoxy group in the present invention and a smaller electronic effect (an effect of providing a molecule with a polarity based on electron-donating and attracting properties) than halogen atom, trifluoromethyl group and methoxy group in the present invention. As a result, the effect of controlling intermolecular interaction in the present invention cannot be expected.

[0053]

Further, compared with 4-, 5-fluorine substituted iridium-phenylpyrimidine complex described

in "Proper Preprints", 2000, 41 (1), 770 - 771, it has been clarified that a luminescence device using the metal coordination compound according to the present invention has a higher durability as shown in Examples appearing later.

[0054]

In the case of phosphorescent (luminescent) material, luminescent characteristics are largely affected by its molecular environment. Principal characteristics of the fluorescent material are studied based on photoluminescence in the case of fluorescence device. However, results of photoluminescence (of the phosphorescent material) do not reflect luminescent characteristics of the resultant EL device in many cases since the luminescent characteristics in the case of the phosphorescence depend on a magnitude of polarity of ambient host material molecules, ambient temperature, whether presence state of the material is solid state or liquid state, etc. Accordingly, it is generally difficult to expect the resultant EL characteristics (for the phosphorescent material) by simply removing a part of characteristics from photoluminescence results.

[0055]

Further, the present inventors have made various studies and found that an organic electro-

luminescence device employing the metal coordination compound represented by the above-mentioned formula (2) (hereinafter, referred to as a "material (2)" in some cases) as the luminescence center material, produces a high-efficient luminescence, exhibits a high luminance (brightness) and provides less deterioration in energized state.

[0056]

The metal coordination compound represented by the above-mentioned formula (2) may preferably have hydrogen atom(s) as at least one of X1 and X2.

[0057]

The metal coordination compound being the material (2) produces phosphorescence (luminescence) and is assumed to have a lowest excited state which is metal-to-ligand charge transfer (MLCT\*) state or  $\pi-\pi^*$  excited state as a triplet excited state. The phosphorescent emission of light (phosphorescence) is produced during the transition from these states to the ground state.

[0058]

A phosphorescence yield and a phosphorescence life can be obtained through a luminescence test based on photoluminescence by photo-excitation. The material (2) has been found to provide a higher phosphorescence yield of 0.15 - 0.9 and a shorter phosphorescence life of 1 - 40  $\mu$ sec. The shorter

phosphorescence life is a condition for providing a resultant EL device with a higher luminescence efficiency. More specifically, the longer phosphorescence life increases molecules placed in their triplet excited state which is a waiting state for phosphorescence, thus resulting in a problem of lowering in luminescence efficiency particularly at a higher current density. The material (2) is a suitable luminescent material for an EL device with a higher phosphorescence yield and a shorter phosphorescence life. Further, by appropriately modifying the alkylene group Y or the substituents X1 and X2 of the metal coordination compound represented by the above-mentioned formula (2), it is possible to expect that emission wavelength can be controlled. From also the viewpoint as described above, the material (2) is a suitable luminescent material for the EL device with a higher phosphorescence yield and a shorter phosphorescence life. Incidentally, the phosphorescence yield may be obtained in the same manner as described with respect to the material (1).  
[0059]

Further, as shown in Examples appearing hereinafter, it has been clarified that the material (2) has also an excellent stability in a continuous energization test. This may be attributable to introduction of particular alkylene group or

substituents allowing control of intermolecular interaction with a host luminescent material and suppression of formation of associated exciton leading to thermal inactivation, thus minimizing quenching to improve device characteristics.

[0060]

The luminescence device according to the present invention, as shown in Figure 1, may preferably include the organic layer comprising the metal coordination compound being the material (1) or the material (2) between two oppositely disposed electrodes comprising a transparent electrode between which are supplied with a voltage to produce luminescence, thus constituting an electric-field luminescence device.

[0061]

The high-efficiency luminescence device shown in the present invention may be applicable to products required to allow energy saving and high luminance, applied examples of which may include those for display apparatus and illumination apparatus, a light source for printers, and backlight (unit) for a liquid crystal display apparatus. As the display apparatus, it is possible to provide a flat panel display apparatus capable of exhibiting an excellent energy saving performance, a high visibility and a good lightweight property. With respect to the light

source for a printer, it becomes possible to replace a laser light source of laser beam printer currently used widely with the luminescence device according to the present invention. Further, when independently addressable devices are arranged in arrays and desired exposure of light to a photosensitive drum is effected to form an image. By the use of the device of the present invention, it becomes possible to considerably reducing the volume (size) of image forming apparatus. With respect to the illumination apparatus and backlight (unit), the resultant apparatus (unit) using the luminescence device of the present invention is expected to have an energy saving effect.

[0062]

<Practice Embodiment 1>

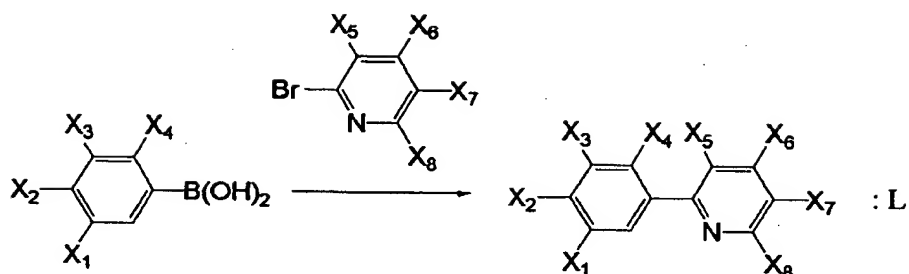
Practice Embodiment 1 relates to the material (1) and a luminescence device thereof.

A synthesis scheme of the metal coordination compound represented by the above-mentioned formula (1) used in the present invention will be shown by taking an iridium complex as an example.

[0063]

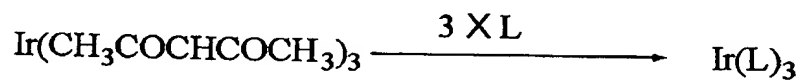
Synthesis of ligand L

[0064]



[0065]

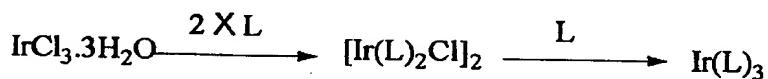
Synthesis of iridium coordination compound



[0067]

or

[0068]



[0069]

Rh coordination compound may be synthesized in the same manner as in the iridium coordination compound.

[0070]

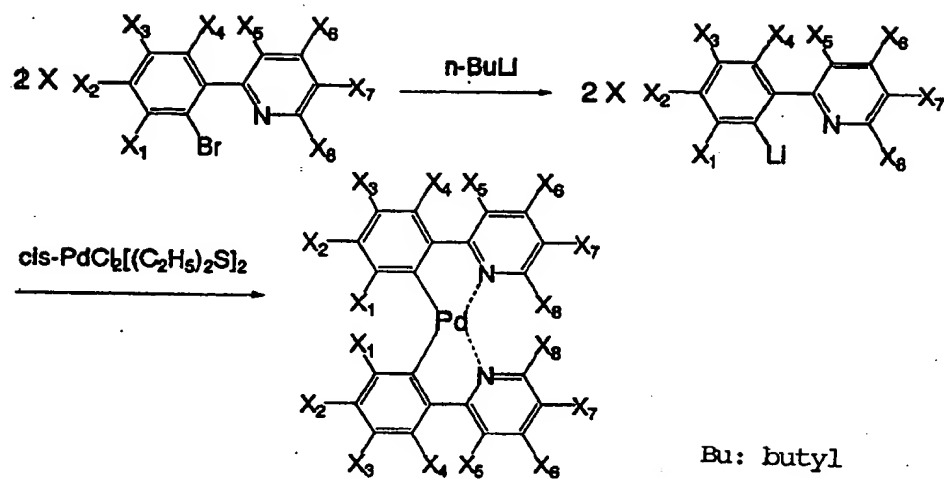
Pd coordination compound is exemplified hereinbelow.

[0071]

Synthesis method of Pd coordination compound.

[0072]





[0073]

Hereinafter, specific structural formulas of the metal coordination compound used in the present invention are shown below. However, these are merely exemplified as representative examples. The present invention is not restricted to these examples.

[0074]

[Table 3]

No.	M	n	X <sub>1</sub>	X <sub>2</sub>	X <sub>3</sub>	X <sub>4</sub>	X <sub>5</sub>	X <sub>6</sub>	X <sub>7</sub>	X <sub>8</sub>
(1)	Ir	3	H	H	F	H	H	H	H	H
(2)	Ir	3	H	F	H	H	H	H	H	H
(3)	Ir	3	H	H	Cl	H	H	H	H	H
(4)	Ir	3	H	H	F	H	H	OCH <sub>3</sub>	H	H
(5)	Ir	3	H	H	F	H	H	H	Br	H
(6)	Ir	3	H	C <sub>2</sub> H <sub>5</sub>	H	H	H	H	H	H
(7)	Ir	3	H	H	NO <sub>2</sub>	H	H	H	H	H
(8)	Ir	3	H	H	NO <sub>2</sub>	H	H	H	CF <sub>3</sub>	H
(9)	Ir	3	H	H	NO <sub>2</sub>	H	H	NO <sub>2</sub>	H	H
(10)	Ir	3	H	H	NO <sub>2</sub>	H	H	OC <sub>11</sub> H <sub>23</sub>	H	H
(11)	Ir	3	H	H	C <sub>3</sub> H <sub>7</sub>	H	H	H	H	H
(12)	Ir	3	H	C <sub>2</sub> H <sub>5</sub>	OCH <sub>3</sub>	H	H	H	H	H
(13)	Ir	3	H	H	C <sub>3</sub> H <sub>7</sub>	H	H	OC <sub>4</sub> H <sub>9</sub>	H	H
(14)	Ir	3	H	C <sub>20</sub> H <sub>41</sub>	H	H	H	H	H	H
(15)	Ir	3	H	H	OCH <sub>3</sub>	H	H	H	H	H
(16)	Ir	3	H	OCH <sub>3</sub>	OCH <sub>3</sub>	H	H	H	H	H
(17)	Ir	3	H	H	OCH(CH <sub>3</sub> ) <sub>2</sub>	H	H	H	H	H
(18)	Ir	3	H	H	OC <sub>6</sub> H <sub>11</sub>	H	H	H	H	H
(19)	Ir	3	H	H	OC <sub>16</sub> H <sub>33</sub>	H	H	H	H	H
(20)	Ir	3	H	H	OCH <sub>3</sub>	H	H	OCH <sub>3</sub>	H	H
(21)	Ir	3	H	H	OCH(CH <sub>3</sub> ) <sub>2</sub>	H	H	OCH <sub>3</sub>	H	H
(22)	Ir	3	H	H	OC <sub>10</sub> H <sub>21</sub>	H	H	NO <sub>2</sub>	H	H
(23)	Ir	3	H	H	OCH(CH <sub>3</sub> ) <sub>2</sub>	H	H	H	CF <sub>3</sub>	H
(24)	Ir	3	H	H	SCH <sub>3</sub>	H	H	H	H	H
(25)	Ir	3	H	OCH <sub>2</sub> CH=CH <sub>2</sub>	H	H	H	H	H	H
(26)	Ir	3	H	H	OCH <sub>2</sub> C≡CCH <sub>3</sub>	H	H	H	H	H
(27)	Ir	3	H	H	COCH <sub>3</sub>	H	H	H	H	H
(28)	Ir	3	H	H	COCH <sub>3</sub>	H	H	NO <sub>2</sub>	H	H
(29)	Ir	3	H	H	COCH <sub>3</sub>	H	H	H	CF <sub>3</sub>	H
(30)	Ir	3	H	H	COCH <sub>3</sub>	H	H	OCH <sub>3</sub>	H	H

[0075]

[Table 4]

No.	M	n	X <sub>1</sub>	X <sub>2</sub>	X <sub>3</sub>	X <sub>4</sub>	X <sub>5</sub>	X <sub>6</sub>	X <sub>7</sub>	X <sub>8</sub>
(31)	Ir	3	H	H	COC <sub>8</sub> H <sub>19</sub>	H	H	H	H	H
(32)	Ir	3	H	H	CF <sub>3</sub>	H	H	H	H	H
(33)	Ir	3	H	H	CF <sub>3</sub>	H	H	H	CF <sub>3</sub>	H
(34)	Ir	3	H	H	CF <sub>3</sub>	H	H	NO <sub>2</sub>	H	H
(35)	Ir	3	H	H	CF <sub>3</sub>	H	H	OCH(CH <sub>3</sub> ) <sub>2</sub>	H	H
(36)	Ir	3	H	C <sub>3</sub> F <sub>7</sub>	H	H	H	H	H	H
(37)	Ir	3	H	H	OCF <sub>3</sub>	H	H	H	H	H
(38)	Ir	3	H	OCF <sub>3</sub>	H	H	H	H	H	H
(39)	Ir	3	H	H	OCF <sub>3</sub>	H	H	NO <sub>2</sub>	H	H
(40)	Ir	3	H	H	OCF <sub>3</sub>	H	H	H	CF <sub>3</sub>	H
(41)	Ir	3	H	H	OCF <sub>3</sub>	H	H	OCH <sub>3</sub>	H	H
(42)	Ir	3	H	H	OCH <sub>2</sub> C <sub>3</sub> F <sub>7</sub>	H	H	H	H	H
(43)	Ir	3	H	O(CH <sub>2</sub> ) <sub>3</sub> C <sub>2</sub> F <sub>6</sub>	H	H	H	H	H	H
(44)	Ir	3	H	H	O(CH <sub>2</sub> ) <sub>3</sub> OCH <sub>2</sub> C <sub>2</sub> F <sub>5</sub>	H	H	H	H	H
(45)	Ir	3	H	H	COOC <sub>2</sub> H <sub>5</sub>	H	H	H	H	H
(46)	Ir	3	H	OCOCH <sub>3</sub>	H	H	H	H	H	H
(47)	Ir	3	H	H	O(CH <sub>2</sub> ) <sub>2</sub> C <sub>3</sub> F <sub>7</sub>	H	H	H	C <sub>6</sub> F <sub>11</sub>	H
(48)	Ir	3	H	H	H	H	H	OCH <sub>3</sub>	H	H
(49)	Ir	3	H	H	H	H	H	H	CF <sub>3</sub>	H
(50)	Ir	3	H	H	H	H	H	H	NO <sub>2</sub>	H
(51)	Ir	3	H	H	Si(CH <sub>3</sub> ) <sub>3</sub>	H	H	H	H	H
(52)	Ir	3	H	H	Si(CH <sub>3</sub> ) <sub>2</sub> C <sub>4</sub> H <sub>9</sub>	H	H	H	H	H
(53)	Ir	3	H	Si(CH <sub>3</sub> ) <sub>2</sub> C <sub>8</sub> H <sub>17</sub>	H	H	H	H	H	H
(54)	Ir	3	H	H	Si(C <sub>2</sub> H <sub>5</sub> ) <sub>3</sub>	H	H	H	H	H
(55)	Ir	3	H	H	H	H	H	Si(CH <sub>3</sub> ) <sub>2</sub> C <sub>6</sub> H <sub>13</sub>	H	H
(56)	Ir	3	H	C <sub>2</sub> H <sub>5</sub>	OCH <sub>3</sub>	H	H	OCH <sub>3</sub>	H	H
(57)	Ir	3	H	F	H	F	H	OCH <sub>3</sub>	H	H
(58)	Ir	3	H	F	H	F	H	OCH <sub>3</sub>	CF <sub>3</sub>	H
(59)	Ir	3	H	H	Si(CH <sub>3</sub> ) <sub>3</sub>	H	H	H	Br	H
(60)	Ir	3	H	Si(CH <sub>3</sub> ) <sub>2</sub> C <sub>7</sub> H <sub>15</sub>	OCH <sub>3</sub>	H	H	H	H	H

[0076]

[Table 5]

No.	M	n	X <sub>1</sub>	X <sub>2</sub>	X <sub>3</sub>	X <sub>4</sub>	X <sub>5</sub>	X <sub>6</sub>	X <sub>7</sub>	X <sub>8</sub>
(61)	Rh	3	H	H	F	H	H	H	H	H
(62)	Rh	3	F	F	H	H	H	H	H	H
(63)	Rh	3	H	H	F	H	H	OCH <sub>3</sub>	H	H
(64)	Rh	3	H	H	NO <sub>2</sub>	H	H	H	H	H
(65)	Rh	3	H	H	NO <sub>2</sub>	H	H	OC <sub>8</sub> H <sub>17</sub>	H	H
(66)	Rh	3	H	H	C <sub>2</sub> H <sub>6</sub>	H	H	H	H	H
(67)	Rh	3	H	C <sub>2</sub> H <sub>6</sub>	OCH <sub>3</sub>	H	H	H	H	H
(68)	Rh	3	H	C <sub>12</sub> H <sub>25</sub>	H	H	H	H	H	H
(69)	Rh	3	H	C <sub>3</sub> H <sub>7</sub>	H	H	H	OCH <sub>3</sub>	H	H
(70)	Rh	3	H	H	OCH(CH <sub>3</sub> ) <sub>2</sub>	H	H	H	H	H
(71)	Rh	3	H	H	OC <sub>15</sub> H <sub>31</sub>	H	H	H	H	H
(72)	Rh	3	H	H	OC <sub>6</sub> H <sub>13</sub>	H	H	NO <sub>2</sub>	H	H
(73)	Rh	3	H	H	OCH <sub>3</sub>	H	H	OCH <sub>3</sub>	H	H
(74)	Rh	3	H	H	OCH(CH <sub>3</sub> ) <sub>2</sub>	H	H	H	CF <sub>3</sub>	H
(75)	Rh	3	H	H	OCH <sub>2</sub> CH=CH <sub>2</sub>	H	H	H	H	H
(76)	Rh	3	H	OC≡CC <sub>4</sub> H <sub>9</sub>	H	H	H	H	H	H
(77)	Rh	3	H	H	SC <sub>2</sub> H <sub>5</sub>	H	H	H	H	H
(78)	Rh	3	H	H	SCH <sub>3</sub>	H	H	OCH <sub>3</sub>	H	H
(79)	Rh	3	H	SCH <sub>3</sub>	SCH <sub>3</sub>	H	H	H	H	H
(80)	Rh	3	H	H	COCH <sub>3</sub>	H	H	H	H	H
(81)	Rh	3	H	H	COCH <sub>3</sub>	H	H	OCH <sub>3</sub>	H	H
(82)	Rh	3	H	H	CF <sub>3</sub>	H	H	H	H	H
(83)	Rh	3	H	H	CF <sub>3</sub>	H	H	OCH(CH <sub>3</sub> ) <sub>2</sub>	H	H
(84)	Rh	3	H	H	OCF <sub>3</sub>	H	H	H	CF <sub>3</sub>	H
(85)	Rh	3	H	H	OCH <sub>2</sub> C <sub>4</sub> F <sub>9</sub>	H	H	H	H	H
(86)	Rh	3	H	H	O(CH <sub>2</sub> ) <sub>6</sub> C <sub>2</sub> F <sub>5</sub>	H	H	H	H	H
(87)	Rh	3	H	H	H	H	H	OCH <sub>3</sub>	H	H
(88)	Rh	3	H	H	Si(CH <sub>3</sub> ) <sub>3</sub>	H	H	H	H	H
(89)	Rh	3	H	Si(CH <sub>3</sub> ) <sub>2</sub> C <sub>6</sub> H <sub>13</sub>	H	H	H	H	H	H
(90)	Rh	3	H	Si(CH <sub>3</sub> ) <sub>2</sub> C <sub>7</sub> H <sub>15</sub>	OCH <sub>3</sub>	H	H	H	H	H

[0077]

[Table 6]

No.	M	n	X <sub>1</sub>	X <sub>2</sub>	X <sub>3</sub>	X <sub>4</sub>	X <sub>5</sub>	X <sub>6</sub>	X <sub>7</sub>	X <sub>8</sub>
(91)	Pd	2	H	H	F	H	H	H	H	H
(92)	Pd	2	H	F	H	F	H	H	H	H
(93)	Pd	2	H	H	F	H	H	OC <sub>7</sub> H <sub>15</sub>	H	H
(94)	Pd	2	H	H	NO <sub>2</sub>	H	H	H	H	H
(95)	Pd	2	H	H	NO <sub>2</sub>	H	H	OC <sub>5</sub> H <sub>11</sub>	H	H
(96)	Pd	2	H	C <sub>2</sub> H <sub>5</sub>	OCH <sub>3</sub>	H	H	H	H	H
(97)	Pd	2	H	H	C <sub>6</sub> H <sub>11</sub>	H	H	OCH <sub>3</sub>	H	H
(98)	Pd	2	H	C <sub>10</sub> H <sub>31</sub>	H	H	H	H	H	H
(99)	Pd	2	H	H	OCH(CH <sub>3</sub> ) <sub>2</sub>	H	H	H	H	H
(100)	Pd	2	H	H	OC <sub>3</sub> H <sub>7</sub>	H	H	H	H	H
(101)	Pd	2	H	H	COC <sub>8</sub> H <sub>17</sub>	H	H	H	H	H
(102)	Pd	2	H	H	CF <sub>3</sub>	H	H	H	H	H
(103)	Pd	2	H	H	CF <sub>3</sub>	H	H	OCH(CH <sub>3</sub> ) <sub>2</sub>	H	H
(104)	Pd	2	H	H	OCF <sub>3</sub>	H	H	H	CF <sub>3</sub>	H
(105)	Pd	2	H	H	Si(CH <sub>3</sub> ) <sub>3</sub>	H	H	H	H	H
(106)	Pd	2	H	H	F	H	H	OC <sub>5</sub> H <sub>11</sub>	H	H
(107)	Pd	2	H	H	NO <sub>2</sub>	H	H	OC <sub>3</sub> H <sub>7</sub>	H	H
(108)	Pd	2	H	H	C <sub>2</sub> H <sub>5</sub>	H	H	OCH <sub>3</sub>	H	H
(109)	Pd	2	H	C <sub>10</sub> H <sub>21</sub>	H	H	H	H	H	H
(110)	Pd	2	H	H	COCH <sub>3</sub>	H	H	H	H	H
(111)	Ir	3	H	Cl	CF <sub>3</sub>	H	H	H	H	H
(112)	Ir	3	H	Cl	CF <sub>3</sub>	H	H	H	CF <sub>3</sub>	H
(113)	Ir	3	H	Cl	CF <sub>3</sub>	H	H	OCH <sub>3</sub>	H	H
(114)	Rh	3	H	Cl	CF <sub>3</sub>	H	H	H	H	H
(115)	Rh	3	H	Cl	CF <sub>3</sub>	H	H	H	CF <sub>3</sub>	H
(116)	Rh	3	H	Cl	CF <sub>3</sub>	H	H	CF <sub>3</sub>	H	H
(117)	Rh	3	H	Cl	CF <sub>3</sub>	H	H	OCH <sub>3</sub>	H	H
(118)	Rh	3	H	Cl	CF <sub>3</sub>	H	H	C <sub>2</sub> H <sub>5</sub>	H	H
(119)	Pd	2	H	Cl	CF <sub>3</sub>	H	H	H	H	H
(120)	Pd	2	H	Cl	CF <sub>3</sub>	H	H	H	CF <sub>3</sub>	CF <sub>3</sub>

[0078]

[Table 7]

No.	M	n	X <sub>1</sub>	X <sub>2</sub>	X <sub>3</sub>	X <sub>4</sub>	X <sub>5</sub>	X <sub>6</sub>	X <sub>7</sub>	X <sub>8</sub>
(121)	Ir	3	H	F	CF <sub>3</sub>	H	H	H	H	H
(122)	Ir	3	H	F	H	F	H	H	H	H
(123)	Ir	3	H	CF <sub>3</sub>	H	CF <sub>3</sub>	H	H	H	H
(124)	Ir	3	H	CF <sub>3</sub>	H	F	H	H	H	H
(125)	Ir	3	H	CF <sub>3</sub>	CF <sub>3</sub>	H	H	H	Br	H
(126)	Ir	3	F	C <sub>2</sub> H <sub>5</sub>	H	H	H	H	H	H
(127)	Ir	3	F	H	NO <sub>2</sub>	H	H	H	H	H
(128)	Ir	3	F	H	NO <sub>2</sub>	F	H	H	CF <sub>3</sub>	H
(129)	Ir	3	F	H	NO <sub>2</sub>	H	H	NO <sub>2</sub>	H	H
(130)	Ir	3	F	H	NO <sub>2</sub>	H	H	OC <sub>11</sub> H <sub>23</sub>	H	H
(131)	Ir	3	F	H	C <sub>3</sub> H <sub>7</sub>	H	H	H	H	H
(132)	Ir	3	F	C <sub>2</sub> H <sub>5</sub>	OCH <sub>3</sub>	H	H	H	H	H
(133)	Ir	4	F	H	C <sub>3</sub> H <sub>7</sub>	H	H	OC <sub>4</sub> H <sub>9</sub>	H	H
(134)	Ir	3	H	C <sub>20</sub> H <sub>41</sub>	H	F	H	H	H	H
(135)	Ir	3	H	H	OCH <sub>3</sub>	F	H	H	H	H
(136)	Ir	3	H	OCH <sub>3</sub>	OCH <sub>3</sub>	F	H	H	H	H
(137)	Ir	3	H	H	OCH(CH <sub>3</sub> ) <sub>2</sub>	F	H	H	H	H
(138)	Ir	3	H	H	OC <sub>8</sub> H <sub>11</sub>	F	H	H	H	H
(139)	Ir	3	H	H	OC <sub>16</sub> H <sub>33</sub>	F	H	H	H	H
(140)	Ir	3	H	H	OCH <sub>3</sub>	F	H	OCH <sub>3</sub>	H	H
(141)	Ir	3	H	H	OCH(CH <sub>3</sub> ) <sub>2</sub>	H	F	OCH <sub>3</sub>	H	H
(142)	Ir	3	H	H	OC <sub>10</sub> H <sub>21</sub>	H	F	NO <sub>2</sub>	H	H
(143)	Ir	3	H	H	OCH(CH <sub>3</sub> ) <sub>2</sub>	H	F	H	CF <sub>3</sub>	H
(144)	Ir	3	H	H	SCH <sub>3</sub>	H	C <sub>2</sub> H <sub>5</sub>	H	H	H
(145)	Ir	3	H	OCH <sub>2</sub> CH=CH <sub>2</sub>	H	H	C <sub>2</sub> H <sub>5</sub>	H	H	H
(146)	Ir	3	H	H	OCH <sub>2</sub> C≡CCH <sub>3</sub>	H	H	H	H	F
(147)	Ir	3	H	H	COCH <sub>3</sub>	H	H	H	H	F
(148)	Ir	3	H	H	COCH <sub>3</sub>	H	H	NO <sub>2</sub>	H	F
(149)	Ir	3	H	H	COCH <sub>3</sub>	H	H	H	CF <sub>3</sub>	F
(150)	Ir	3	CF <sub>3</sub>	H	COCH <sub>3</sub>	H	H	OCH <sub>3</sub>	H	H

[0079]

[Table 8]

No.	M	n	X <sub>1</sub>	X <sub>2</sub>	X <sub>3</sub>	X <sub>4</sub>	X <sub>5</sub>	X <sub>6</sub>	X <sub>7</sub>	X <sub>8</sub>
(151)	Ir	3	F	H	COC <sub>6</sub> H <sub>19</sub>	H	H	H	H	H
(152)	Ir	3	H	CF <sub>3</sub>	H	F	H	H	H	H
(153)	Ir	3	F	H	CF <sub>3</sub>	H	H	H	CF <sub>3</sub>	H
(154)	Ir	3	H	H	CF <sub>3</sub>	F	H	NO <sub>2</sub>	H	H
(155)	Ir	3	H	H	OF <sub>3</sub>	F	H	OCH(CH <sub>3</sub> ) <sub>2</sub>	H	H
(156)	Ir	3	H	C <sub>3</sub> F <sub>7</sub>	H	CF <sub>3</sub>	H	H	H	H
(157)	Ir	3	H	H	OCF <sub>3</sub>	H	CF <sub>3</sub>	H	H	H
(158)	Ir	3	H	OCF <sub>3</sub>	H	H	C <sub>2</sub> H <sub>5</sub>	H	H	H
(159)	Ir	3	H	CF <sub>3</sub>	H	CF <sub>3</sub>	H	H	H	H
(160)	Ir	3	H	H	OCF <sub>3</sub>	H	F	H	CF <sub>3</sub>	H
(161)	Ir	3	H	H	OCF <sub>3</sub>	H	H	OCH <sub>3</sub>	H	F
(162)	Ir	3	H	H	OCH <sub>2</sub> C <sub>3</sub> F <sub>7</sub>	H	H	H	H	F
(163)	Ir	3	H	O(CH <sub>2</sub> ) <sub>3</sub> C <sub>2</sub> F <sub>5</sub>	H	H	H	H	H	F
(164)	Ir	3	H	H	O(CH <sub>2</sub> ) <sub>3</sub> OCH <sub>2</sub> C <sub>2</sub> F <sub>5</sub>	Cl	H	H	H	H
(165)	Ir	3	H	H	COOC <sub>2</sub> H <sub>5</sub>	F	H	H	H	H
(166)	Rh	3	H	OCOCH <sub>3</sub>	H	H	F	H	H	H
(167)	Rh	3	H	H	O(CH <sub>2</sub> ) <sub>2</sub> C <sub>3</sub> F <sub>7</sub>	H	C <sub>2</sub> H <sub>5</sub>	H	C <sub>6</sub> F <sub>11</sub>	H
(168)	Rh	3	H	H	H	H	H	OCH <sub>3</sub>	H	F
(169)	Rh	3	H	H	H	H	H	H	CF <sub>3</sub>	F
(170)	Rh	3	H	H	H	H	H	H	NO <sub>2</sub>	F
(171)	Rh	3	H	H	Si(CH <sub>3</sub> ) <sub>3</sub>	H	H	H	H	F
(172)	Rh	3	H	H	Si(CH <sub>3</sub> ) <sub>2</sub> C <sub>4</sub> H <sub>9</sub>	H	H	H	H	F
(173)	Rh	3	H	Si(CH <sub>3</sub> ) <sub>2</sub> C <sub>6</sub> H <sub>17</sub>	H	F	H	H	H	H
(174)	Rh	3	H	H	Si(C <sub>2</sub> H <sub>5</sub> ) <sub>3</sub>	F	H	H	H	H
(175)	Rh	3	H	H	H	F	H	SK(CH <sub>3</sub> ) <sub>2</sub> C <sub>6</sub> H <sub>13</sub>	H	H
(176)	Pd	2	H	C <sub>2</sub> H <sub>5</sub>	OCH <sub>3</sub>	F	H	OCH <sub>3</sub>	H	H
(177)	Pd	2	F	H	F	F	H	OCH <sub>3</sub>	H	H
(178)	Pd	2	F	H	F	F	H	OCH <sub>3</sub>	CF <sub>3</sub>	H
(179)	Pd	2	F	H	Si(CH <sub>3</sub> ) <sub>3</sub>	H	H	H	Br	H
(180)	Pd	2	F	Si(CH <sub>3</sub> ) <sub>2</sub> C <sub>7</sub> H <sub>15</sub>	OCH <sub>3</sub>	H	H	H	H	H

[0080]

<Examples 1 - 20>

Examples 1 - 20 relates to the material (1) and luminescence device using the material.

[0081]

Examples 1-10 and Comparative Example 1

A common portion of device preparation steps used in the present invention will be described.

[0082]

As a device structure, a device having a three-layer structure of organic layers shown in (b) of Figure 1 was used. On a glass substrate (transparent substrate 15), a 100 nm-thick film (transparent electrode 14) of ITO (indium tin oxide) was formed by sputtering, followed by patterning to have an (opposing) electrode area of 3 mm<sup>2</sup>. On the ITO-formed substrate, organic layers and metal electrode layers shown below were successively formed in film by vacuum (vapor) deposition using resistance heating in a vacuum chamber (10<sup>-4</sup> Pa).

Organic layer 1 (hole transport layer 13) (40 nm):  $\alpha$ -NPD

Organic layer 2 (luminescence layer 12) (30 nm): mixture of CBP:metal coordination compound (metal coordination compound weight proportion: 5 wt. %)

Organic layer 3 (electron transport layer 16) (30 nm): Alq3



Metal electrode layer 1 (15 nm): Al-Li alloy  
(Li = 1.8 wt. %)

Metal electrode layer 2 (100 nm): Al  
[0083]

An electric field was applied between the ITO electrode (as an anode) and the Al electrode (as a cathode) to apply a voltage to each device so as to provide the device with the same current value, thus measuring a luminance (brightness) with time. The constant current amount was 70 mA/cm<sup>2</sup>. At that time, the resultant luminances of the respective devices were in the range of 70 - 210 cd/m<sup>2</sup>.

[0084]

As the cause of device deterioration, oxygen or moisture (water content) is a problematic factor, so that each device was subjected to the above measurement in a dry nitrogen gas stream after it was taken out of the vacuum chamber so as to remove such a factor.

[0085]

In Comparative Example 1, as the conventional luminescent material, Ir(ppy)<sub>3</sub> described in the aforementioned document 2.

[0086]

[Table 9]

Ex. No.	Luminescence material No.	Luminance half-life (Hr)
1	(4)	750
2	(7)	500
3	(17)	900
4	(18)	850
5	(21)	850
6	(23)	500
7	(32)	600
8	(56)	700
9	(67)	400
10	(74)	450
Comp. Ex. 1	Ir(ppy) <sub>3</sub>	350

[0087]

The results of energization durability test of the devices using the respective compounds are shown in Table 9. The devices (of the present invention) provide luminance half-life times clearly longer than the device using the conventional luminescent material, so that it becomes possible to realize a high-durability device resulting from stability of the materials of the present invention.

[0088]

Examples 11 - 13 and Comparative Example 2

In these examples, a device having a four-layer structure of organic layers shown in (c) of Figure 1 was used.

On a glass substrate (transparent substrate 15), a 100 nm-thick film (transparent electrode 14) of ITO (indium tin oxide) was formed, followed by patterning to have an (opposing) electrode area of  $3 \text{ mm}^2$ . On the ITO-formed substrate, organic layers and metal electrode layers shown below were successively formed in film by vacuum (vapor) deposition using resistance heating in a vacuum chamber ( $10^{-4}$  Pa).

Organic layer 1 (hole transport layer 13) (40 nm):  $\alpha$ -NPD

Organic layer 2 (luminescence layer 12) (20 nm): mixture of CBP: predetermined coordination compound (weight proportion: 7 wt. %)

Organic layer 3 (exciton diffusion prevention layer 17) (10 nm): BCP

Organic layer 4 (electron transport layer 16) (30 nm): Alq3

Metal electrode layer 1 (15 nm): Al-Li alloy (Li = 1.8 wt. %)

Metal electrode layer 2 (100 nm): Al  
[0089]

As the metal coordination compounds, Ex. Comp. Nos. (1), (32) and (49) were used. These

compounds were subjected to measurement of photoluminescence spectrum in order to evaluate a luminescent characteristic of these metal coordination compounds. Each of the metal coordination compounds in a solution state in toluene at a concentration of  $10^{-4}$  mol/l and at 25 °C was subjected to measurement of photoluminescence spectrum by using a spectrophoto-fluorometer ("Model F4500", mfd. by Hitachi K.K.). An excited light of approximately 350 nm was used. The values of photoluminescence spectrum of the metal coordination compounds were substantially equivalent to those in the EL devices under voltage application, so that it was confirmed that luminescence produced by the EL device was based on luminescence from the metal coordination compound used.

[0090]

Experimental results of the above photoluminescence and EL devices were shown in Table 10.

[0091]

Characteristics of the EL devices were measured by using a microammeter ("Model 4140B", mfd. by Hewlett-Packard Co.) for current-voltage characteristic (current), using a spectrophotofluorometer ("Model SR1", mfd. by Topcon K.K.) for luminescence spectrum (peak emission wavelength

$\lambda_{PE}$ ), and using a luminance meter ("Model BM7", mfd. by Topcon K.K.) for a luminescence luminance (luminescence efficiency). Further, an energy conversion efficiency (lm/W) was a ratio obtained according to the following equation:

$$\begin{aligned} \text{Energy conversion efficiency (lm/W)} \\ = (\pi \times \text{luminescence efficiency} \\ (\text{cd/A})) / \text{applied voltage (V)}. \end{aligned}$$

In these examples, the respective devices corresponding to the respective coordination compounds showed a good electrical rectification characteristic. [0092]

Although the luminescence device using  $\text{Ir(ppy)}_3$  (Comparative Example 2) shows a maximum (peak emission wavelength) of emission spectrum ( $\lambda_{PE}$ ) of 510 nm, the EL devices using the metal coordination compounds used in the present invention are found to show longer peak emission wavelengths than  $\text{Ir(ppy)}_3$  by approximately 15 - 35 nm, thus resulting in smaller relative luminous efficiencies.

Although smaller energy conversion efficiencies and luminescence efficiencies of the luminescence devices of the present invention are obtained compared with those of the luminescence device using  $\text{Ir(ppy)}_3$ , this may be attributable to the smaller relative luminous efficiencies with the longer peak emission wavelengths, thus not attributable to

essentially inferior luminescent characteristics of the luminescence devices using the metal coordination compounds of these Examples.

[0093]

With respect to the luminance half-lives of the luminescence devices, experiment was performed in the same manner as in Examples 1 - 11. Compared with the luminescence device using  $\text{Ir(ppy)}_3$ , the luminescence devices using the metal coordination compounds of these Examples are found to show longer luminance half-lives.

[0094]

[Table 10]

Ex.No	Luminescent material No. (Ex. Comp. No.)	$\lambda$ PE in toluene (nm)	$\lambda$ PE (nm)	Energy conversion efficiency (lm/W)	Luminescence efficiency (cd/A)	Current density (mA/cm <sup>2</sup> at 12V)	Luminance half-life (Hr)
Comp. Ex. 2	Ir(ppy) <sub>3</sub>	510	510	6.0	19.0	20	150
11	(1)	522	525	4.0	13.6	170	300
12	(32)	487	525	0.4	2.4	130	400
13	(49)	537	545	2.1	7.0	25	250

[0095]

Synthesis examples of the metal coordination compounds used in the present invention will be shown below.

[0096]

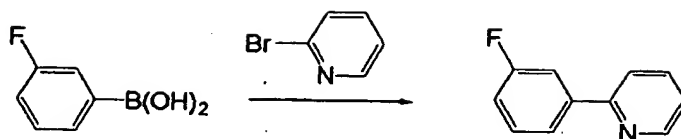
Example 14 (Synthesis of Ex. Comp. No. (1))

In a 1 liter-three necked flask, 20.0 g (126.6 mM) of 2-bromopyridine, 17.7 g (126.4 mM) of 3-fluorophenylbromic acid, 130 ml of toluene, 65 ml of ethanol and 130 ml of 2M-sodium carbonate aqueous solution were placed and stirred in a nitrogen gas stream at room temperature. Under stirring, to the mixture, 4.60 g (3.98 mM) of tetrakis (triphenylphosphine) palladium (0) was added, followed by heat-refluxing for 6 hours under stirring in nitrogen gas stream.

After the reaction, the reaction mixture was cooled, followed by extraction with cool water and toluene. The organic layer was washed with water until the system showed neutral, followed by distilling off of the solvent under reduced pressure to obtain a residue. The residue was purified by silica gel column chromatography (eluent: toluene/ethyl acetate = 5/1) to obtain 6.0 g of 2-(3-fluorophenyl)pyridine (pale brown liquid) (Yield: 34.6 %).

[0097]



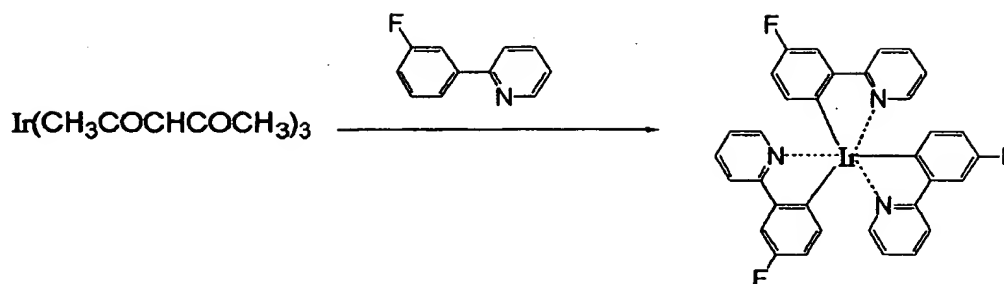


[0098]

In a 100 ml-four necked flask, 50 ml of glycerol was placed and heat-stirred for 2 hours at 130 - 140 °C while supplying nitrogen gas therein in the form of bubbles, followed by cooling to 100 °C by standing. To glycerol, 1.04 g (6.00 mM) of 2-(3-fluorophenyl)pyridine and 0.50 g (1.02 mM) of Iridium (III) acetylacetonate were added, followed by heat-refluxing for 10 hours under stirring in nitrogen gas stream.

After the reaction, the reaction mixture was cooled to room temperature and poured into 300 ml of 1N-HCl. The resultant precipitate was recovered by filtration and washed with water, followed by drying for 5 hours at 100 °C under reduced pressure and purification by silica gel column chromatography (eluent: chloroform) to obtain 0.22 g of Iridium (III) tris[2-(3-fluorophenyl)pyridine] (yellow powder) (Yield: 31.0 %).

[0099]



[0100]

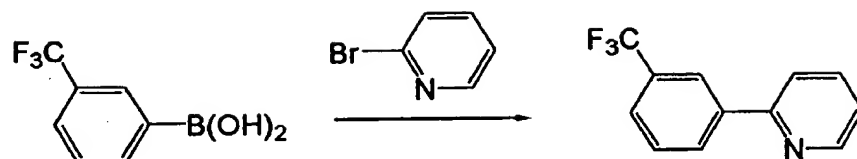
**Example 15** (Synthesis of Ex. Comp. No. (32))

In a 1 liter-three necked flask, 20.8 g (131.6 mM) of 2-bromopyridine, 25.0 g (131.6 mM) of 3-trifluoromethylphenylbromic acid, 130 ml of toluene, 65 ml of ethanol and 130 ml of 2M-sodium carbonate aqueous solution were placed and stirred in a nitrogen gas stream at room temperature. Under stirring, to the mixture, 4.76 g (4.12 mM) of tetrakis (triphenylphosphine) palladium (0) was added, followed by heat-refluxing for 7 hours under stirring in nitrogen gas stream.

After the reaction, the reaction mixture was cooled, followed by extraction with cool water and toluene. The organic layer was washed with water until the system showed neutral, followed by distilling off of the solvent under reduced pressure to obtain a residue (pale brown liquid). The residue was purified by silica gel column chromatography (eluent: toluene/hexane = 1/1) to obtain 6.0 g of 2-(3-trifluoromethylphenyl)pyridine (pale brown liquid)

(Yield: 21.1%).

[0101]



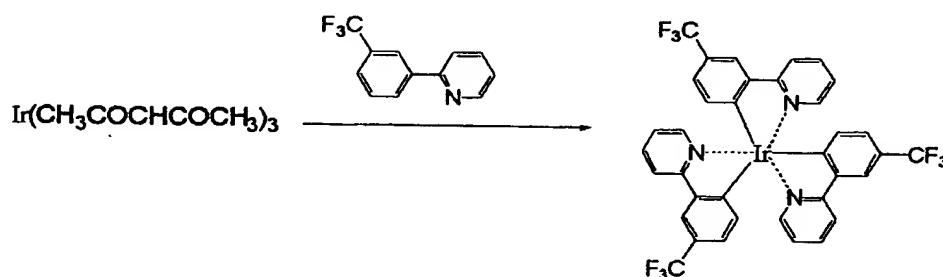
[0102]

In a 200 ml-four necked flask, 100 ml of glycerol was placed and heat-stirred for 2 hours at 130 - 140 °C while supplying nitrogen gas therein in the form of bubbles, followed by cooling to 100 °C by standing. To glycerol, 2.68 g (12.0 mM) of 2-(3-trifluoromethylphenyl)pyridine and 1.00 g (2.04 mM) of Iridium (III) acetylacetonate were added, followed by heat-refluxing for 10 hours under stirring in nitrogen gas stream.

After the reaction, the reaction mixture was cooled to room temperature and poured into 600 ml of 1N-HCl. The resultant precipitate was recovered by filtration and washed with water, followed by drying for 5 hours at 100 °C under reduced pressure. The precipitate was dissolved in chloroform and the insoluble matter was removed by filtration, followed by purification by silica gel column chromatography (eluent: chloroform) and recrystallization from a

mixture solvent (chloroform/methanol) to obtain 0.62 g of Iridium (III) tris[2-(3-trifluoromethylphenyl)-pyridine] (yellow powder) (Yield: 35.3 %) showed a peak emission wavelength  $\lambda_{\text{max}}$  of emission spectrum in toluene at 25 °C of 487 nm.

[0103]



[0104]

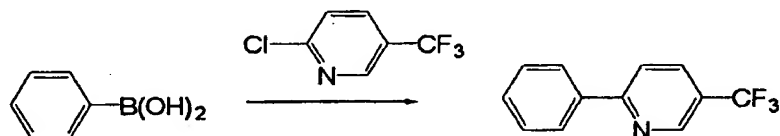
Example 16 (Synthesis of Ex. Comp. No. (49))

In a 1 liter-three necked flask, 25.6 g (141.0 mM) of 2-chloro-5-trifluoromethylpyridine, 17.2 g (141.0 mM) of phenylboronic acid, 140 ml of toluene, 70 ml of ethanol and 140 ml of 2M-sodium carbonate aqueous solution were placed and stirred in a nitrogen gas stream at room temperature. Under stirring, to the mixture, 5.10 g (4.41 mM) of tetrakis (triphenylphosphine) palladium (0) was added, followed by heat-refluxing for 6 hours under stirring in nitrogen gas stream.

After the reaction, the reaction mixture was cooled, followed by extraction with cool water and

toluene. The organic layer was washed with water until the system showed neutral, followed by distilling off of the solvent under reduced pressure to obtain a residue. The residue was purified by silica gel column chromatography (eluent: toluene/hexane = 5/1). The resultant creamy crystal was purified by alumina column chromatography (eluent: toluene) and recrystallized from ethanol to obtain 13.1 g of 2-phenyl-5-trifluoromethylpyridine (colorless crystal) (Yield: 41.6 %).

[0105]



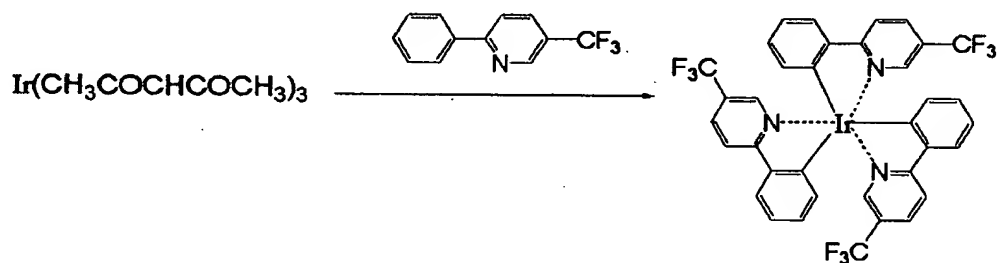
[0106]

In a 200 ml-four necked flask, 100 ml of glycerol was placed and heat-stirred for 2 hours at 130 - 140 °C while supplying nitrogen gas therein in the form of bubbles, followed by cooling to 100 °C by standing. To glycerol, 2.68 g (12.0 mM) of 2-phenyl-5-trifluoromethylpyridine and 1.00 g (2.04 mM) of Iridium (III) acetylacetonate were added, followed by heat-refluxing for 8 hours under stirring in nitrogen gas stream.

After the reaction, the reaction mixture was cooled to room temperature and poured into 600 ml of

1N-HCl. The resultant precipitate was recovered by filtration and washed with water, followed by drying for 4 hours at 100 °C under reduced pressure and purification by silica gel column chromatography (eluent: chloroform) to obtain 0.43 g of Iridium (III) tris-(2-phenyl-5-trifluoromethylpyridine) (orange powder) (Yield: 24.5 %).

[0107]



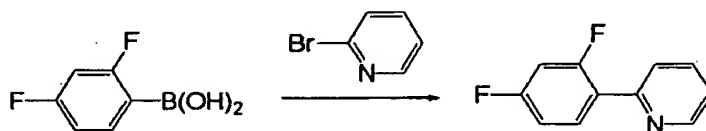
[0108]

Example 17 (Synthesis of Ex. Comp. No. (122))

In a 100 ml-three necked flask, 3.16 g (19.9 mM) of 2-bromopyridine, 3.16 g (20.0 mM) of 2,4-difluorophenylboronic acid, 15 ml of toluene, 7.5 ml of ethanol and 15 ml of 2M-sodium carbonate aqueous solution were placed and stirred in a nitrogen gas stream at room temperature. Under stirring, to the mixture, 0.72 g (0.62 mM) of tetrakis (triphenylphosphine) palladium (0) was added, followed by heat-refluxing for 8 hours and 40 minutes under stirring in nitrogen gas stream.

After the reaction, the reaction mixture was cooled, followed by extraction with cool water and ethyl acetate. The organic layer was washed with water followed by distilling off of the solvent under reduced pressure to obtain a residue. The residue was purified by silica gel column chromatography (eluent: toluene/ethyl acetate = 10/1) to obtain 3.28 g of 2-(2,4-difluorophenyl)pyridine (pale yellow oily product) (Yield: 86.0 %).

[0109]



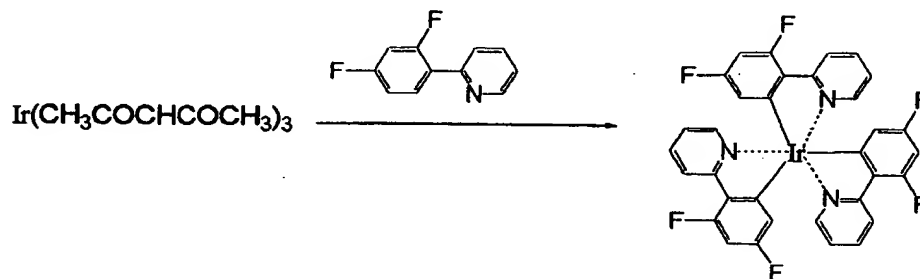
[0110]

In a 100 ml-four necked flask, 50 ml of glycerol was placed and heat-stirred for 2 hours at 130 - 140 °C while supplying nitrogen gas therein in the form of bubbles, followed by cooling to 100 °C by standing. To glycerol, 0.96 g (5.02 mM) of 2-(2,4-difluorophenyl)pyridine and 0.50 g (1.02 mM) of Iridium (III) acetylacetonate were added, followed by heat-refluxing for 10 hours under stirring in nitrogen gas stream.

After the reaction, the reaction mixture was cooled to room temperature and poured into 300 ml of 1N-HCl. The resultant precipitate was recovered by

filtration and washed with water, followed by drying for 5 hours at 100 °C under reduced pressure and purification by silica gel column chromatography (eluent: chloroform) and recrystallization from a mixture solvent (chloroform/methanol) to obtain 0.25 g of Iridium (III) tris[2-(4,6-difluorophenyl)-pyridine] (yellow powder) (Yield: 32.1 %), which showed a peak emission wavelength  $\lambda_{\text{max}}$  of emission spectrum in toluene at 25 °C of 471 nm.

[0111]



[0112]

Example 18 (Synthesis of Ex. Comp. No. (121))

In a 500 ml-three necked flask, 11.0 g (45.3 mM) of 5-bromo-2-fluorobenzotrifluoride and 90 ml of dry tetrahydrofuran (THF) were placed and stirred in a nitrogen gas stream at room temperature. Under stirring, to the mixture, 2.60 g (2.25 mM) of tetrakis(triphenylphosphine) palladium (0) was added, followed by cooling to 20 - 21 °C (inner temperature) on an ice bath in nitrogen gas stream. At that

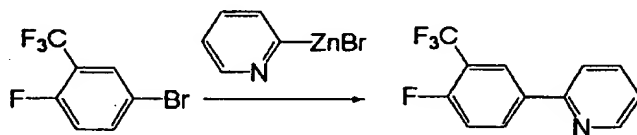


temperature, 90 ml of 0.5 M-THF solution of 2-pyridylzinc bromide was gradually added dropwise to the mixture in nitrogen gas stream, followed by stirring for 4 hours at that temperature.

After the reaction, the reaction mixture was poured into cool water, followed by addition of ethyl acetate to remove the insoluble matter by filtration. The organic layer was washed with water and dried with anhydrous sodium sulfate, followed by distilling-off of the solvent under reduced pressure to obtain a residue.

The residue was purified by silica gel column chromatography (eluent: hexane/ethyl acetate = 20/1) to obtain 1.80 g of 2-(4-fluoro-3-trifluoromethylphenyl)pyridine (pale brown oily product) (Yield: 16.6 %).

[0113]



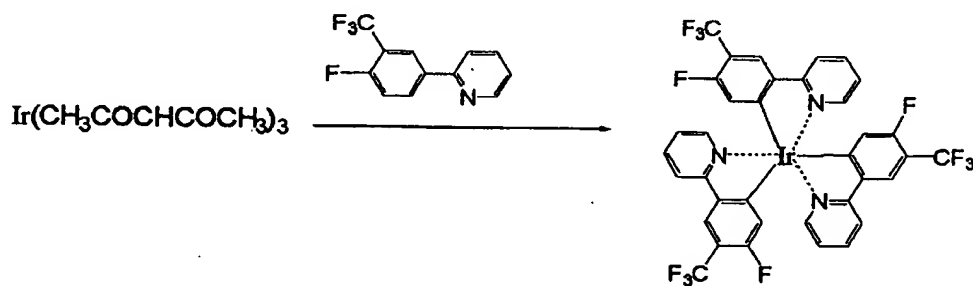
[0114]

In a 100 ml-four necked flask, 50 ml of glycerol was placed and heat-stirred for 2 hours at 130 - 140 °C while supplying nitrogen gas therein in the form of bubbles, followed by cooling to 100 °C by standing. To glycerol, 1.21 g (5.02 mM) of 2-(4-

fluoro-3-trifluoromethylphenyl)pyridine and 0.50 g (1.02 mM) of Iridium (III) acetylacetonate were added, followed by heat-refluxing for 10 hours under stirring in nitrogen gas stream.

After the reaction, the reaction mixture was cooled to room temperature and poured into 300 ml of 1N-HCl. The resultant precipitate was recovered by filtration and washed with water, followed by drying for 5 hours at 100 °C under reduced pressure and purification by silica gel column chromatography (eluent: chloroform) and recrystallization from a mixture solvent (chloroform/methanol) to obtain 0.20 g of Iridium (III) tris[2-(4-fluoro-5-trifluoromethylphenyl)pyridine] (yellow powder) (Yield: 21.5 %), which showed a peak emission wavelength  $\lambda_{\text{max}}$  of emission spectrum in toluene at 25 °C of 466 nm.

[0115]



[00116]

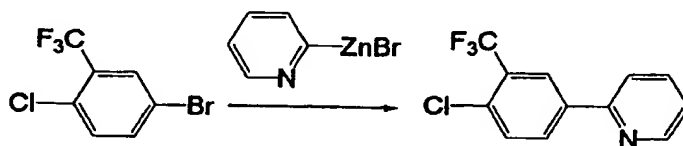
Example 19 (Synthesis of Ex. Comp. No. (111))

In a 500 ml-three necked flask, 11.8 g (45.5 mM) of 5-bromo-2-chlorobenzotrifluoride and 90 ml of

dry tetrahydrofuran (THF) were placed and stirred in a nitrogen gas stream at room temperature. Under stirring, to the mixture, 2.60 g (2.25 mM) of tetrakis(triphenylphosphine) palladium (0) was added, followed by cooling to 13.5 - 14 °C (inner temperature) on an ice bath in nitrogen gas stream. At that temperature, 90 ml of 0.5 M-THF solution of 2-pyridylzinc bromide was gradually added dropwise to the mixture in nitrogen gas stream, followed by stirring for 3 hours at ca. 20 °C.

After the reaction, the reaction mixture was poured into cool water, followed by addition of ethyl acetate to remove the insoluble matter by filtration. The organic layer was washed with water and dried with anhydrous sodium sulfate, followed by distilling-off of the solvent under reduced pressure to obtain a residue. The residue was purified by silica gel column chromatography (eluent: hexane/ethyl acetate = 10/1) to obtain 3.70 g of 2-(4-chloro-5-trifluoromethylphenyl)pyridine (pale brown oily product) (Yield: 31.9 %).

[0117]

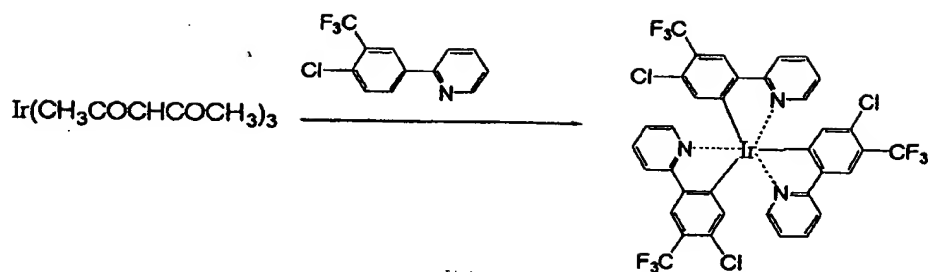


[0118]

In a 100 ml-four necked flask, 50 ml of glycerol was placed and heat-stirred for 2 hours at 130 - 140 °C while supplying nitrogen gas therein in the form of bubbles, followed by cooling to 100 °C by standing. To glycerol, 1.29 g (5.01 mM) of 2-(4-chloro-3-trifluoromethylphenyl)pyridine and 0.50 g (1.02 mM) of Iridium (III) acetylacetonate were added, followed by heat-refluxing for 8 hours under stirring in nitrogen gas stream.

After the reaction, the reaction mixture was cooled to room temperature and poured into 300 ml of 1N-HCl. The resultant precipitate was recovered by filtration and washed with water, followed by drying for 5 hours at 100 °C under reduced pressure and purification by silica gel column chromatography (eluent: chloroform) and recrystallization from a mixture solvent (chloroform/hexane) to obtain 0.25 g of Iridium (III) tris[2-(4-chloro-3-trifluoromethylphenyl)pyridine] (yellow powder) (Yield: 25.4 %), which showed a peak emission wavelength  $\lambda_{\text{max}}$  of emission spectrum in toluene at 25 °C of 479 nm.

[0119]



[0120]

Comparative Example 3 (Synthesis of metal coordination compound A)

Iridium (III) tris[2-(4,5-difluoromethylphenyl)pyridine metal coordination compound A described in Polymer Preprints, 41(1), pp. 770 - 771 (2000) was prepared in the same manner as in Example 17 except that 2,4-difluorophenylbionic acid was changed to 3,4-difluorophenylbionic acid.

The metal coordination compound A showed a peak emission wavelength  $\lambda_{\text{max}}$  of emission spectrum in toluene at 25 °C of 505 nm.

[0121]

Example 20 and Comparative Example 4

Two luminescence devices were prepared and subjected to continuous energization test in the same manner as in Example 1 by using the metal coordination compound (Ex. Comp. No. (122)) and the metal coordination compound A prepared in Comparative Example 3, respectively.

The results are shown in Table 11 below.

[0122]

[Table 11]

Ex. No.	Luminescent material No. (Ex. Comp. No.)	Luminance half-life (Hr)
20	(122)	630
Comp.Ex. 4	Metal coordination compound A	310

[0123]

As apparent from Table 18, the luminescence device using the metal coordination compound of formula (1) according to the present invention exhibited a luminance half-life considerably longer than that of the luminescence device using the metal coordination compound A, thus resulting in an EL device excellent in durability (luminance stability).

[0124]

As described hereinabove, the metal coordination compound used in the present invention (material (1)) provides a higher phosphorescence efficiency and a shorter phosphorescence life and allows control of its emission wavelength by appropriately modifying the substituents X1 to X8, thus being suitable as a luminescent material for EL device.

[0125]

The luminescence device having an organic layer containing the metal coordination compound is an excellent device which exhibits a high efficiency luminescence, a high luminance for a long period, and a less luminescence deterioration in energized state.

<Practice Embodiment 2>

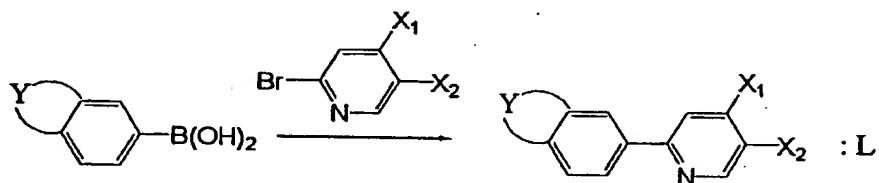
Embodiment 2 relates to the material (2) and the luminescence device using the material.

A synthesis scheme of the metal coordination compound represented by the above-mentioned formula (2) will be shown by taking the iridium compound as an example.

[0127]

Synthesis of ligand L

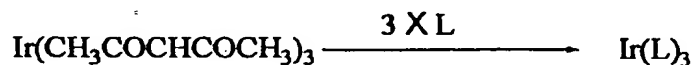
[0128]



[0129]

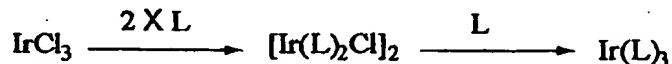
Synthesis of iridium coordination compound

[0130]



or

[0131]



[0132]

Rh coordination compound may be synthesized in the same manner as in iridium coordination compound shown above.

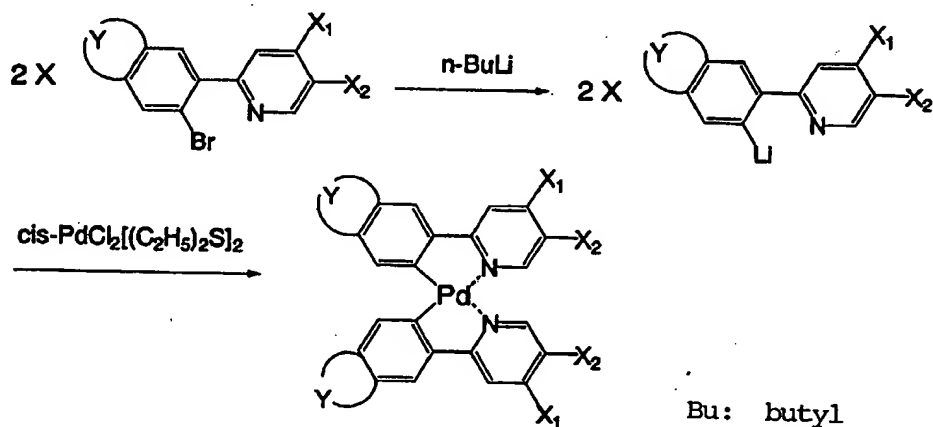
[0133]

Pd coordination compound is exemplified below.

[0134]

Synthesis method of Pd coordination compound

[0135]

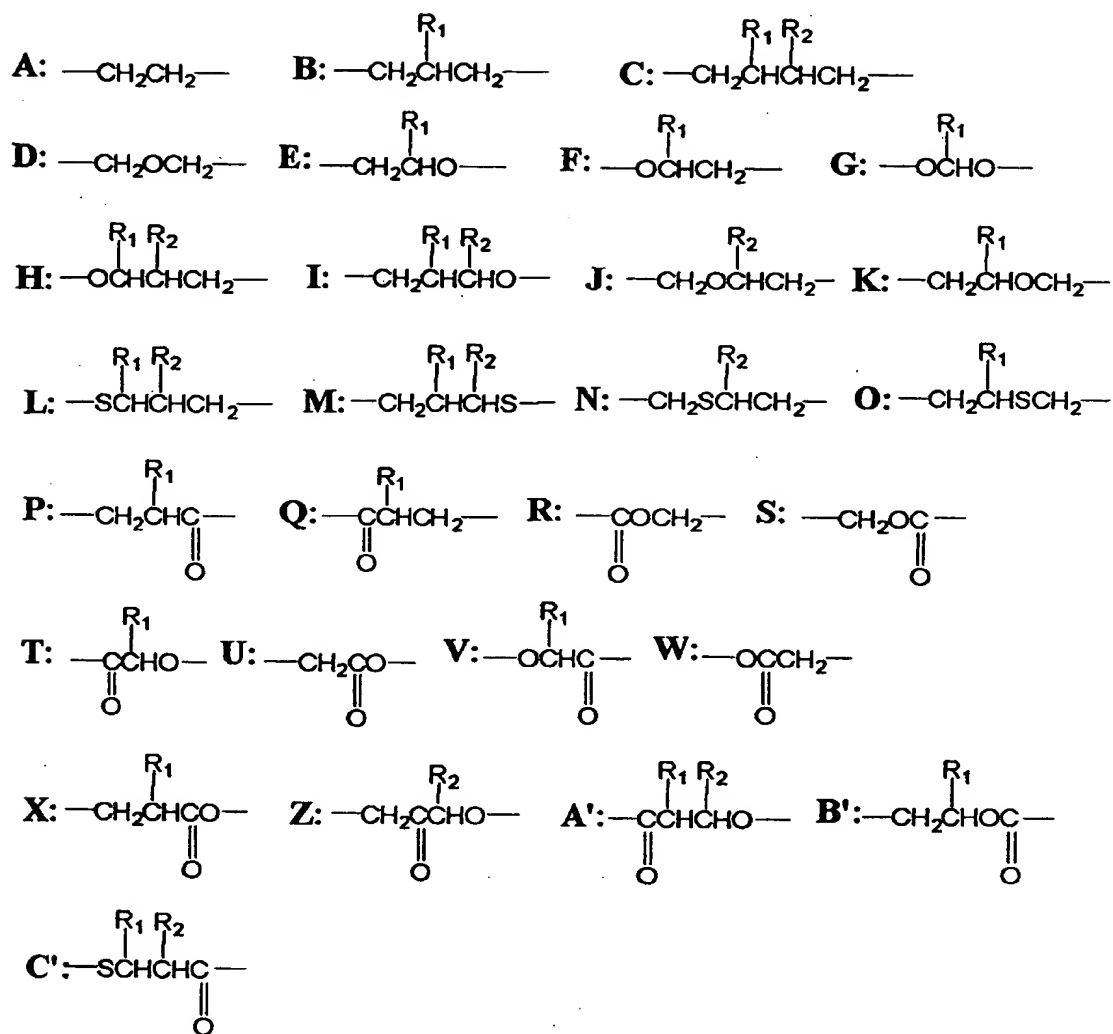


[0136]

Hereinbelow, specific structural formulas of the metal coordination compounds used in the present invention are shown in Tables 12 - 18. However, there are merely exemplified as representative examples. The present invention is not limited thereto. A to C used in Tables 12 - 18 represent formulas shown below.



[0137]



[0138]

[Table 12]

No.	M	n	Y	R <sub>1</sub>	R <sub>2</sub>	X <sub>1</sub>	X <sub>2</sub>
2001	Ir	3	A	—	—	H	H
2002	Ir	3	A	—	—	OOH <sub>3</sub>	H
2003	Ir	3	B	H	—	H	H
2004	Ir	3	B	H	—	OCH <sub>3</sub>	H
2005	Ir	3	B	H	—	H	OF <sub>3</sub>
2006	Ir	3	B	H	—	H	Cl
2007	Ir	3	B	CH <sub>3</sub>	—	H	H
2008	Ir	3	B	CH <sub>3</sub>	—	F	H
2009	Ir	3	B	CH <sub>3</sub>	—	NO <sub>2</sub>	H
2010	Ir	3	B	C <sub>2</sub> H <sub>5</sub>	—	H	H
2011	Ir	3	B	C <sub>3</sub> H <sub>7</sub>	—	H	CF <sub>3</sub>
2012	Ir	3	B	C <sub>2</sub> H <sub>5</sub> (CH <sub>2</sub> )CHCH <sub>2</sub>	—	H	H
2013	Ir	4	B	C <sub>6</sub> H <sub>13</sub>	—	OCH(CH <sub>3</sub> ) <sub>2</sub>	H
2014	Ir	3	B	C <sub>10</sub> H <sub>21</sub>	—	Si(CH <sub>3</sub> ) <sub>3</sub>	H
2015	Ir	3	C	H	H	H	H
2016	Ir	3	C	H	H	OCH <sub>3</sub>	H
2017	Ir	3	C	H	H	H	OF <sub>3</sub>
2018	Ir	3	C	H	H	F	H
2019	Ir	3	C	H	H	NO <sub>2</sub>	H
2020	Ir	3	C	H	H	OC <sub>3</sub> H <sub>11</sub>	H
2021	Ir	3	C	H	H	O(CH <sub>2</sub> ) <sub>2</sub> C <sub>3</sub> F <sub>7</sub>	H
2022	Ir	3	C	H	H	H	Si(C <sub>2</sub> H <sub>5</sub> ) <sub>3</sub>
2023	Ir	3	C	H	H	H	Br
2024	Ir	3	C	H	H	CH <sub>3</sub>	H
2025	Ir	3	C	CH <sub>3</sub>	H	CH <sub>3</sub>	H
2026	Ir	3	C	H	CH <sub>3</sub>	H	H
2027	Ir	3	C	CH <sub>3</sub>	CH <sub>3</sub>	H	H
2028	Ir	3	C	C <sub>2</sub> H <sub>5</sub>	H	Si(CH <sub>3</sub> ) <sub>3</sub>	H
2029	Ir	3	C	H	C <sub>2</sub> H <sub>11</sub>	H	H
2030	Ir	3	C	O <sub>6</sub> H <sub>17</sub>	H	Cl	H

[0139]

[Table 13]

No.	M	n	Y	R <sub>1</sub>	R <sub>2</sub>	X <sub>1</sub>	X <sub>2</sub>
2031	Ir	3	C	C <sub>2</sub> H <sub>5</sub>	C <sub>2</sub> H <sub>5</sub>	H	C <sub>7</sub> F <sub>15</sub>
2032	Ir	3	C	H	C <sub>6</sub> H <sub>13</sub>	NO <sub>2</sub>	H
2033	Ir	3	C	C <sub>10</sub> H <sub>21</sub>	H	CF <sub>3</sub>	H
2034	Ir	3	C	H	C <sub>9</sub> H <sub>19</sub>	H	OC <sub>4</sub> H <sub>9</sub>
2035	Ir	3	D	—	—	H	H
2036	Ir	3	D	—	—	OCH <sub>3</sub>	H
2037	Ir	3	E	H	—	H	H
2038	Ir	3	E	H	—	H	NO <sub>2</sub>
2039	Ir	3	E	CH <sub>3</sub>	—	H	H
2040	Ir	3	E	CH <sub>3</sub>	—	OCH <sub>3</sub>	H
2041	Ir	3	E	CH <sub>3</sub>	—	H	CF <sub>3</sub>
2042	Ir	3	E	CH <sub>3</sub>	—	NO <sub>2</sub>	H
2043	Ir	3	E	CH <sub>3</sub>	—	OC <sub>3</sub> H <sub>7</sub>	H
2044	Ir	3	E	C <sub>2</sub> H <sub>5</sub>	—	H	H
2045	Ir	3	E	C <sub>2</sub> H <sub>5</sub>	—	H	CF <sub>3</sub>
2046	Ir	3	E	C <sub>3</sub> H <sub>7</sub>	—	H	H
2047	Ir	3	E	C <sub>3</sub> H <sub>7</sub>	—	OC <sub>5</sub> H <sub>11</sub>	H
2048	Ir	3	E	(CH <sub>2</sub> ) <sub>2</sub> CHCH <sub>2</sub> CH <sub>2</sub>	—	H	H
2049	Ir	3	E	C <sub>5</sub> H <sub>11</sub>	—	H	C <sub>4</sub> F <sub>9</sub>
2050	Ir	3	E	C <sub>5</sub> H <sub>13</sub>	—	H	H
2051	Ir	3	E	C <sub>5</sub> H <sub>13</sub>	—	H	Br
2052	Ir	3	E	C <sub>5</sub> H <sub>13</sub>	—	NO <sub>2</sub>	H
2053	Ir	3	E	C <sub>5</sub> H <sub>17</sub>	—	H	H
2054	Ir	3	E	C <sub>9</sub> H <sub>19</sub>	—	OCH <sub>2</sub> C≡CCH <sub>3</sub>	H
2055	Ir	3	E	C <sub>10</sub> H <sub>21</sub>	—	H	H
2056	Ir	3	E	C <sub>10</sub> H <sub>21</sub>	—	OCH <sub>2</sub> CH=CH <sub>2</sub>	H
2057	Ir	3	F	H	—	OCH <sub>3</sub>	H
2058	Ir	3	F	CH <sub>3</sub>	—	H	H
2059	Ir	3	F	CH <sub>3</sub>	—	OCH <sub>3</sub>	H
2060	Ir	3	F	C <sub>2</sub> H <sub>5</sub>	—	H	CF <sub>3</sub>

[0140]

[Table 14]

No.	M	n	Y	R <sub>1</sub>	R <sub>2</sub>	X <sub>1</sub>	X <sub>2</sub>
2061	Ir	3	F	C <sub>6</sub> H <sub>13</sub>	—	OCH(CH <sub>3</sub> ) <sub>2</sub>	H
2062	Ir	3	F	C <sub>8</sub> H <sub>17</sub>	—	Si(CH <sub>3</sub> ) <sub>2</sub> C <sub>8</sub> H <sub>17</sub>	H
2063	Ir	3	G	H	—	OCH <sub>3</sub>	H
2064	Ir	3	G	H	—	H	CF <sub>3</sub>
2065	Ir	3	G	H	—	O(CH <sub>2</sub> ) <sub>3</sub> OCH <sub>2</sub> C <sub>2</sub> F <sub>5</sub>	H
2066	Ir	3	G	CH <sub>3</sub>	—	H	H
2067	Ir	3	H	H	H	H	H
2068	Ir	3	H	CH <sub>3</sub>	H	Si(CH <sub>3</sub> ) <sub>3</sub>	H
2069	Ir	3	H	H	CH <sub>3</sub>	H	Cl
2070	Ir	3	I	H	H	H	H
2071	Ir	3	I	H	H	OCH <sub>3</sub>	H
2072	Ir	3	I	H	H	H	CF <sub>3</sub>
2073	Ir	3	I	H	H	H	CH <sub>3</sub>
2074	Ir	3	I	C <sub>2</sub> H <sub>5</sub>	H	COOC <sub>2</sub> H <sub>5</sub>	H
2075	Ir	3	I	H	C <sub>6</sub> H <sub>11</sub>	OCH <sub>2</sub> CH=CH <sub>2</sub>	H
2076	Ir	3	J	H	—	H	H
2077	Ir	3	J	H	—	NO <sub>2</sub>	H
2078	Ir	3	J	CH <sub>3</sub>	—	OCH <sub>3</sub>	H
2079	Ir	3	K	H	—	H	H
2080	Ir	3	K	H	—	H	Si(CH <sub>3</sub> ) <sub>3</sub>
2081	Ir	3	K	C <sub>3</sub> H <sub>7</sub>	—	H	CF <sub>3</sub>
2082	Ir	3	L	H	H	H	H
2083	Ir	3	L	CH <sub>3</sub>	H	SC <sub>2</sub> H <sub>5</sub>	H
2084	Ir	3	L	H	CH <sub>3</sub>	OC <sub>6</sub> H <sub>13</sub>	H
2085	Ir	3	M	H	H	H	H
2086	Ir	3	M	C <sub>2</sub> H <sub>5</sub>	H	COOC <sub>2</sub> H <sub>7</sub>	H
2087	Ir	3	M	H	C <sub>2</sub> H <sub>5</sub>	H	O(CH <sub>2</sub> ) <sub>3</sub> C <sub>2</sub> F <sub>5</sub>
2088	Ir	3	N	—	H	H	H
2089	Ir	3	N	—	C <sub>2</sub> H <sub>5</sub>	H	NO <sub>2</sub>
2090	Ir	3	N	—	C <sub>6</sub> H <sub>13</sub>	Cl	H

[0141]

[Table 15]

No.	M	n	Y	R <sub>1</sub>	R <sub>2</sub>	X <sub>1</sub>	X <sub>2</sub>
2091	Ir	3	O	H	—	H	H
2092	Ir	3	O	H	—	H	Si(C <sub>2</sub> H <sub>5</sub> ) <sub>3</sub>
2093	Ir	3	O	C <sub>8</sub> H <sub>17</sub>	—	OCH(CH <sub>3</sub> ) <sub>2</sub>	H
2094	Ir	3	P	H	—	H	H
2095	Ir	3	P	C <sub>3</sub> H <sub>7</sub>	—	H	COOCH <sub>3</sub>
2096	Ir	3	P	C <sub>8</sub> H <sub>19</sub>	—	H	H
2097	Ir	3	Q	H	—	H	H
2098	Ir	3	Q	C <sub>6</sub> H <sub>5</sub>	—	O(CH <sub>2</sub> ) <sub>3</sub> CH=CH <sub>2</sub>	H
2099	Ir	3	R	—	—	H	H
2100	Ir	3	R	—	—	H	CF <sub>3</sub>
2101	Ir	3	S	—	—	H	H
2102	Ir	3	S	—	—	OC <sub>2</sub> H <sub>5</sub>	H
2103	Ir	3	T	H	—	H	Br
2104	Ir	3	T	C <sub>2</sub> H <sub>5</sub>	—	H	H
2105	Ir	3	U	—	—	H	H
2106	Ir	3	U	—	—	H	C <sub>7</sub> F <sub>15</sub>
2107	Ir	3	V	H	—	H	H
2108	Ir	3	W	—	—	OCH <sub>2</sub> C≡CCH <sub>3</sub>	H
2109	Ir	3	X	CH <sub>3</sub>	—	H	H
2110	Ir	3	Z	—	H	O(CH <sub>2</sub> ) <sub>2</sub> CH(CH <sub>3</sub> ) <sub>2</sub>	H
2111	Ir	3	Z	—	C <sub>2</sub> H <sub>7</sub>	H	H
2112	Ir	3	A'	H	H	H	H
2113	Ir	3	B'	H	—	H	NO <sub>2</sub>
2114	Ir	3	B'	CH <sub>3</sub>	—	H	H
2115	Ir	3	C'	H	C <sub>8</sub> H <sub>19</sub>	OCH <sub>3</sub>	H
2116	Pt	2	A	—	—	H	H
2117	Pt	2	B	H	—	H	H
2118	Pt	2	B	H	—	H	C <sub>4</sub> F <sub>9</sub>
2119	Pt	2	B	CH <sub>3</sub>	—	OCH <sub>3</sub>	H
2120	Pt	2	B	C <sub>3</sub> H <sub>7</sub>	—	H	CF <sub>3</sub>

[0142]

[Table 16]

No.	M	n	Y	R <sub>1</sub>	R <sub>2</sub>	X <sub>1</sub>	X <sub>2</sub>
2121	Pt	2	B	C <sub>8</sub> H <sub>17</sub>	—	H	H
2122	Pt	2	C	H	H	H	H
2123	Pt	2	C	H	H	H	CF <sub>3</sub>
2124	Pt	2	C	CH <sub>3</sub>	CH <sub>3</sub>	H	H
2125	Pt	2	C	C <sub>2</sub> H <sub>5</sub>	H	H	H
2126	Pt	2	C	C <sub>10</sub> H <sub>21</sub>	H	OCH <sub>3</sub>	H
2127	Pt	2	D	—	—	H	H
2128	Pt	2	E	H	—	H	H
2129	Pt	2	E	CH <sub>3</sub>	—	H	H
2130	Pt	2	E	CH <sub>3</sub>	—	H	H
2131	Pt	2	E	CH <sub>3</sub>	—	H	NO <sub>2</sub>
2132	Pt	2	E	C <sub>6</sub> H <sub>13</sub>	—	OC <sub>2</sub> H <sub>5</sub>	H
2133	Pt	2	F	CH <sub>3</sub>	—	H	H
2134	Pt	2	F	C <sub>2</sub> H <sub>5</sub>	—	H	OF <sub>3</sub>
2135	Pt	2	G	H	—	H	H
2136	Pt	2	G	H	—	H	Si(CH <sub>3</sub> ) <sub>3</sub>
2137	Pt	2	G	C <sub>4</sub> H <sub>9</sub>	—	H	CH <sub>3</sub>
2138	Pt	2	H	H	C <sub>8</sub> H <sub>19</sub>	H	H
2139	Pt	2	I	H	H	H	H
2140	Pt	2	I	C <sub>2</sub> H <sub>5</sub>	H	H	Si(C <sub>2</sub> H <sub>5</sub> ) <sub>3</sub>
2141	Pt	2	J	—	H	H	H
2142	Pt	2	K	C <sub>8</sub> H <sub>11</sub>	—	H	H
2143	Pt	2	L	C <sub>8</sub> H <sub>17</sub>	H	SO <sub>2</sub> H <sub>3</sub>	H
2144	Pt	2	N	—	H	H	H
2145	Pt	2	O	H	—	H	H
2146	Pt	2	P	H	—	H	H
2147	Pt	2	Q	H	—	H	CH <sub>3</sub>
2148	Pt	2	R	—	—	H	H
2149	Pt	2	U	—	—	H	H
2150	Pt	2	V	H	—	NO <sub>2</sub>	H

[0143]

[Table 17]

No.	M	n	Y	R <sub>1</sub>	R <sub>2</sub>	X <sub>1</sub>	X <sub>2</sub>
2151	Pt	2	W	—	—	H	H
2152	Pt	2	X	CH <sub>3</sub>	—	H	H
2153	Pt	2	Z	—	H	H	H
2154	Pt	2	A'	H	H	H	H
2155	Pt	2	B'	H	—	OCH <sub>3</sub>	H
2156	Pt	2	C'	H	H	H	CF <sub>3</sub>
2157	Rh	3	B	H	—	H	Br
2158	Rh	3	B	H	—	OC <sub>6</sub> H <sub>13</sub>	H
2159	Rh	3	B	CH <sub>3</sub>	—	H	H
2160	Rh	3	C	H	H	H	H
2161	Rh	3	C	H	H	OCH <sub>3</sub>	H
2162	Rh	3	C	H	H	NO <sub>2</sub>	H
2163	Rh	3	C	H	CH <sub>3</sub>	H	H
2164	Rh	3	C	C <sub>6</sub> H <sub>13</sub>	H	H	Si(CH <sub>3</sub> ) <sub>3</sub>
2165	Rh	3	D	—	—	H	H
2166	Rh	3	E	H	—	COOC <sub>2</sub> H <sub>5</sub>	H
2167	Rh	3	E	CH <sub>3</sub>	—	H	H
2168	Rh	3	E	CH <sub>3</sub>	—	H	O(CH <sub>2</sub> ) <sub>6</sub> C <sub>2</sub> F <sub>5</sub>
2169	Rh	3	E	C <sub>3</sub> H <sub>7</sub>	—	H	H
2170	Rh	3	E	C <sub>10</sub> H <sub>21</sub>	—	H	H
2171	Rh	3	F	C <sub>8</sub> H <sub>17</sub>	—	H	H
2172	Rh	3	G	H	—	OCH <sub>2</sub> CH=CH <sub>2</sub>	H
2173	Rh	3	G	CH <sub>3</sub>	—	H	CF <sub>3</sub>
2174	Rh	3	H	H	H	H	H
2175	Rh	3	I	H	H	H	H
2176	Rh	3	K	C <sub>2</sub> H <sub>5</sub>	—	Cl	H
2177	Rh	3	M	H	H	H	H
2178	Rh	3	N	—	H	H	H
2179	Rh	3	P	CH <sub>3</sub>	—	H	NO <sub>2</sub>
2180	Rh	3	S	—	—	H	H

[0144]

[Table 18]

No.	M	n	Y	R <sub>1</sub>	R <sub>2</sub>	X <sub>1</sub>	X <sub>2</sub>
2181	Rh	3	V	H	—	H	H
2182	Rh	3	X	H	—	SC <sub>5</sub> H <sub>11</sub>	H
2183	Rh	3	C'	H	—	OC <sub>7</sub> H <sub>15</sub>	H
2184	Pd	2	B	C <sub>6</sub> H <sub>13</sub>	—	H	H
2185	Pd	2	C	H	H	OCH <sub>3</sub>	H
2186	Pd	2	C	H	H	H	H
2187	Pd	2	D	—	—	H	H
2188	Pd	2	E	H	—	H	CF <sub>3</sub>
2189	Pd	2	E	CH <sub>3</sub>	—	H	H
2190	Pd	2	F	C <sub>3</sub> H <sub>7</sub>	—	H	H
2191	Pd	2	G	H	—	H	H
2192	Pd	2	G	H	—	Si(OH) <sub>3</sub>	H
2193	Pd	2	I	CH <sub>3</sub>	H	NO <sub>2</sub>	H
2194	Pd	2	J	—	H	H	H
2195	Pd	2	L	H	H	H	H
2196	Pd	2	M	H	H	C <sub>4</sub> F <sub>9</sub>	H
2197	Pd	2	O	H	—	H	C <sub>4</sub> H <sub>9</sub>
2198	Pd	2	T	H	—	H	H
2199	Pd	2	W	—	—	OCH <sub>3</sub>	OCH <sub>3</sub>
2200	Pd	2	A'	CH <sub>3</sub>	H	H	Cl



[0145]

Examples 21 - 39 relate to the material (2) and the luminescence device thereof.

[0146]

Examples 21 - 39 and Comparative Example 5

A common portion of device preparation steps used in the present invention will be described.

[0147]

As a device structure, a device having a three-layer structure shown in Figure 1(b) was used. On a glass substrate (transparent substrate 15), a 100 nm-thick film (transparent electrode 14) of ITO (indium tin oxide) was formed, followed by patterning to have an (opposing) electrode area of 3 mm<sup>2</sup>. On the ITO-formed substrate, organic layers and metal electrode layers shown below were successively formed by vacuum (vapor) deposition using resistance heating in a vacuum chamber (10<sup>-4</sup> Pa).

Organic layer 1 (hole transport layer 13) (40 nm):  $\alpha$ -NPD

Organic layer 2 (luminescence layer 12) (20 nm): mixture of CBP: metal coordination compound (metal coordination compound proportion: 5 wt.%)

Organic layer 3 (electron transport layer 16) (30 nm): Alq3

Metal electrode layer 1 (15 nm): Al-Li alloy (Li = 1.8 wt. %)

Metal electrode layer 2 (100 nm): Al

[0148]

An electric field was applied between the ITO electrode (as an anode) and the Al electrode (as a cathode) to apply a voltage to each device so as to provide the device with the same current value, thus measuring a luminance (brightness) with time. The constant current amount was  $70 \text{ mA/cm}^2$ . At that time, the resultant luminances of the respective devices were in the range of  $60 - 220 \text{ cd/m}^2$ .

[0149]

As the cause of device deterioration, oxygen or moisture (water content) is a problematic factor, so that each device was subjected to the above measurement in a dry nitrogen gas stream after it was taken out of the vacuum chamber so as to remove such a factor.

[0150]

In Comparative Example 5, as the conventional luminescent material,  $\text{Ir(ppy)}_3$  described in the aforementioned document 2 (the same as Comparative Example 1).

[0151]

The results of energization durability test of the devices using the respective compounds are shown in Table 19. The devices (of the present invention) provide luminance half-life times clearly

longer than the device using the conventional luminescent material, so that it becomes possible to realize a high-durability device resulting from stability of the materials of the present invention.

[0152]

[Table 19]

Ex. No.	Luminescent material No. (Ex.Comp. No.)	Luminance half-life (Hr)
21	(2010)	750
22	(2015)	950
23	(2017)	800
24	(2021)	850
25	(2039)	900
26	(2043)	750
27	(2046)	900
28	(2085)	500
29	(2096)	650
30	(2122)	650
31	(2131)	600
32	(2146)	550
33	(2163)	600
34	(2177)	450
35	(2182)	450
Comp.Ex. 5	Ir(ppy) <sub>3</sub>	350

[0153]

Examples 36 and 37 and Comparative Example 6

In these examples, as a device structure, a device having a four-layer structure shown in Figure 1(c) was used. On a glass substrate (transparent substrate 15), a 100 nm-thick film (transparent electrode 14) of ITO (indium tin oxide) was formed, followed by patterning to have an (opposing) electrode area of 3 mm<sup>2</sup>. On the ITO-formed substrate, organic layers and metal electrode layers shown below were successively formed by vacuum (vapor) deposition using resistance heating in a vacuum chamber (10<sup>-4</sup> Pa).

Organic layer 1 (hole transport layer 13) (40 nm):  $\alpha$ -NPD

Organic layer 2 (luminescence layer 12) (20 nm): mixture of CBP: predetermined metal coordination compound (weight proportion: 7 wt. %)

Organic layer 3 (exciton diffusion prevention layer 17) (10 nm): BCP

Organic layer 4 (electron transport layer 16) (30 nm): Alq3

Metal electrode layer 1 (15 nm): Al-Li alloy (Li = 1.8 wt. %)

Metal electrode layer 2 (100 nm): Al

[0154]

As the metal coordination compounds, Ex. Comp. Nos. (2015) and (2017) were used. These

compounds were subjected to measurement of photoluminescence spectrum in order to evaluate a luminescent characteristic of these metal coordination compounds. Each of the metal coordination compounds in a solution state in toluene at a concentration of  $10^{-4}$  mol/l and at 25 °C was subjected to measurement of photoluminescence spectrum at by using a spectrophoto-fluorometer ("Model F4500", mfd. by Hitachi K.K.). An excited light of approximately 350 nm was used. The values of photoluminescence spectrum of the metal coordination compounds were substantially equivalent to those in the EL devices under voltage application, so that it was confirmed that luminescence produced by the EL device was based on luminescence from the metal coordination compound used.

[0155]

Experimental results of the above photoluminescence and EL devices were shown in Table 20.

[0156]

Characteristics of the EL devices were measured by using a microammeter ("Model 4140B", mfd. by Hewlett-Packard Co.) for current-voltage characteristic (current), using a spectrophotofluorometer ("Model SR1", mfd. by Topcon K.K.) for luminescence spectrum (peak emission wavelength

$\lambda_{PE}$ ), and using a luminance meter ("Model BM7", mfd. by Topcon K.K.) for a luminescence luminance (luminescence efficiency). In these examples, the respective devices corresponding to the respective coordination compounds showed a good electrical rectification characteristic.

[0157]

Although the luminescence device using  $\text{Ir(ppy)}_3$  (Comparative Example 6 similar to Comparative Example 2) shows a maximum (peak emission wavelength) of emission spectrum ( $\lambda_{PE}$ ) of 510 nm, the EL devices using the metal coordination compounds used in the present invention are found to show longer peak emission wavelengths than  $\text{Ir(ppy)}_3$  by approximately 55 nm, thus resulting in smaller relative luminous efficiencies.

Although smaller energy conversion efficiencies and luminescence efficiencies of the luminescence devices of the present invention are obtained compared with those of the luminescence device using  $\text{Ir(ppy)}_3$ , this may be attributable to the smaller relative luminous efficiencies with the longer peak emission wavelengths, thus not attributable to essentially inferior luminescent characteristics of the luminescence devices using the metal coordination compounds of these Examples.

[0158]

With respect to the luminance half-lives of the luminescence devices, experiment was performed in the same manner as in Examples 1 - 11. Compared with the luminescence device using  $\text{Ir(ppy)}_3$ , the luminescence devices using the metal coordination compounds of these Examples are found to show longer luminance half-lives.

[0159]

[Table 20]

Ex.No	Luminescent material No.	$\lambda$ PE in toluene (nm)	$\lambda$ PE (nm)	Energy conversion efficiency (lm/W)	Luminescence efficiency (cd/A)	Current density (mA/cm <sup>2</sup> at 12V)	Luminance half-life (Hr)
Comp. Ex. 6	Ir(ppy) <sub>3</sub>	510	510	6.0	19.0	20	150
16	2015	524	565	0.9	7.5	70	250
17	2017	554	565	3.4	9.6	180	300



[0160]

Synthesis examples of the metal coordination compounds used in the present invention will be shown below.

[0161]

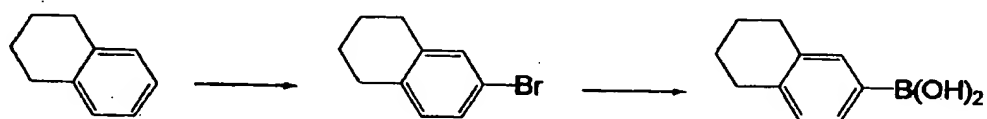
Example 38 (Synthesis of Ex. Comp. No. (2015))

In a 5 liter-three necked flask, 169.5 g (1.28 M) of 1,2,3,4-tetrahydronaphthalene and 3 liters of acetic acid were placed and stirred at room temperature. Under stirring, to the mixture, 650 g (1.67 M) of benzyltrimethylammonium bromide and 244.8 g (1.80 M) of zinc chloride were successively added, followed by stirring for 5.5 hours at 70 °C. After the reaction, the reaction mixture was cooled to room temperature and poured into 3 liters of ice water, followed by extraction with methyl t-butyl ether. The organic layer was successively washed with 5 %-NaHSO<sub>3</sub> aqueous solution, 5 %-NaOH aqueous solution and distilled water, followed by distilling-off of the solvent under reduced pressure to obtain 243.2 g of a dark brown liquid. The liquid was subjected to vacuum distillation (distillation under reduced pressure) (boiling point = 108 - 110 °C at 667 Pa) to obtain 130.2 g of 6-bromo-1,2,3,4-tetrahydronaphthalene (Yield: 48.1 %).

In a 5 liter-three necked flask, 67.55 g of 6-bromo-1,2,3,4-tetrahydronaphthalene and 1480 ml of

dry tetrahydrofuran (THF) were placed and cooled to -70 to -68 °C on a dry ice-acetone bath in a dry nitrogen gas atmosphere. At that temperature, to the mixture, 200 ml of 1.6 M-butyllithium solution in hexane was added dropwise, followed by stirring for 2 hours at -67 °C or below. To the mixture, 69.8 g (672 mM) of trimethyl borate in 435 ml of dry THF was added dropwise at -70 to -68 °C, followed by stirring for 2 hours at -67 °C or below. The reaction mixture was gradually warmed to room temperature and left standing overnight. The resultant reaction mixture was gradually added dropwise to a mixture of 108 ml of HCl and 438 ml of water kept at 10 °C or below, followed by stirring for 1 hour at that temperature. Thereafter, the mixture was subjected to extraction with toluene. The organic layer was washed with water, followed by distilling-off of the solvent under reduced pressure to obtain a residue. The residue was purified by silica gel column chromatography (eluent: toluene/ethyl acetate = 2/1) and recrystallized from hexane to obtain 30.4 g of 1,2,3,4-tetrahydronaphthalene-6-boronic acid (Yield: 54.0 %).

[0162]

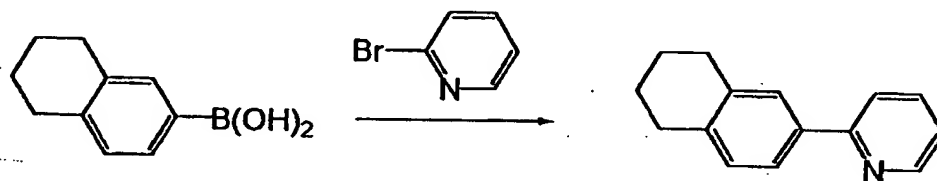


[0163]

In a 1 liter-three necked flask, 17.8 g (114 mM) of 2-bromopyridine, 20.0 g (127 mM) of 1,2,3,4-tetrahydronaphthalene-6-bionic acid, 160 ml of toluene, 80 ml of ethanol and 160 ml of 2M-sodium carbonate aqueous solution were placed and stirred in a nitrogen gas stream at room temperature. Under stirring, to the mixture, 4.05 g (3.5 mM) of tetrakis (triphenylphosphine) palladium (0) was added, followed by heat-refluxing for 7 hours under stirring in nitrogen gas stream.

After the reaction, the reaction mixture was cooled, followed by extraction with cool water and toluene, and distilling off of the solvent under reduced pressure to obtain a residue. The residue was purified by silica gel column chromatography (eluent: toluene/hexane = 2/1) to obtain 9.2 g of 6-(pyridine-2-yl)-1,2,3,4-tetrahydronaphthalene (yellow liquid) (Yield: 38.6 %).

[0164]



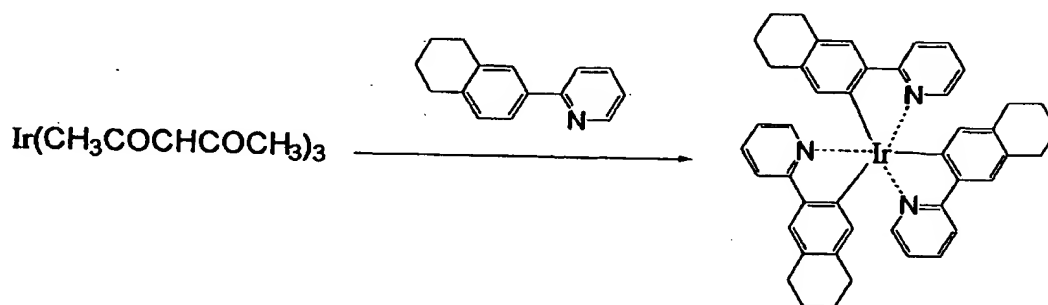
[0165]

In a 100 ml-four necked flask, 50 ml of

glycerol was placed and heat-stirred for 2 hours at 130 - 140 °C while supplying nitrogen gas therein in the form of bubbles, followed by cooling to 100 °C by standing. To glycerol, 1.30 g (6.21 mM) of 6-(pyridine-2-yl)-1,2,3,4-tetrahydronaphthalene and 0.50 g (1.02 mM) of Iridium (III) acetylacetonate were added, followed by heat-refluxing for 5 hours under stirring in nitrogen gas stream.

After the reaction, the reaction mixture was cooled to room temperature and poured into 100 ml of 1N-HCl. The resultant precipitate was recovered by filtration and washed with water, followed by washing with acetone and purification by silica gel column chromatography (eluent: chloroform) to obtain 0.14 g of Iridium (III) tris[6-(pyridine-2-yl)-1,2,3,4-tetrahydronaphthalene] (orange powder) (Yield: 16.8 %).

[0166]



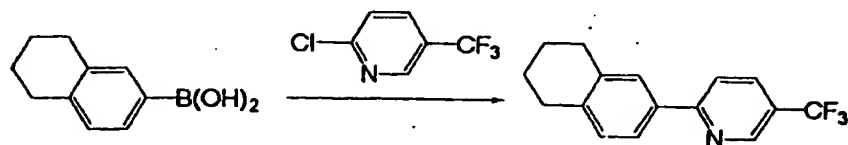
[0167]

Example 39 (Synthesis of Ex. Comp. No. (2017))

In a 200 ml-four necked flask, 5.16 g (28.4 mM) of 2-chloro-5-trifluoromethyl, 5.00 g (28.4 mM) of 1,2,3,4-tetrahydronaphthalene-6-bionic acid, 25 ml of toluene, 12.5 ml of ethanol and 25 ml of 2M-sodium carbonate aqueous solution were placed and stirred in a nitrogen gas stream at room temperature. Under stirring, to the mixture, 1.02 g (0.88 mM) of tetrakis (triphenylphosphine) palladium (0) was added, followed by heat-refluxing for 3.25 hours under stirring in nitrogen gas stream.

After the reaction, the reaction mixture was cooled, followed by extraction with cool water and toluene, and distilling off of the solvent under reduced pressure to obtain a residue. The residue was purified by silica gel column chromatography (eluent: toluene/hexane = 1/1) and alumina column chromatography (eluent: toluene) and recrystallized from methanol to obtain 3.14 g of 6-(5-trifluoromethylpyridine-2-yl)-1,2,3,4-tetrahydronaphthalene (colorless crystal) (Yield: 39.9 %).

[0168]

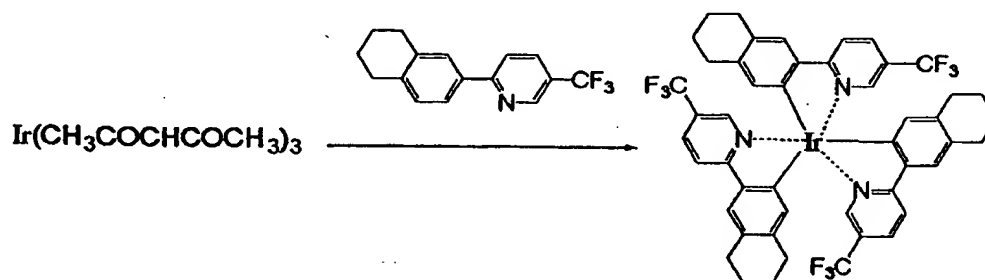


[0169]

In a 100 ml-four necked flask, 50 ml of glycerol was placed and heat-stirred for 2 hours at 130 - 140 °C while supplying nitrogen gas therein in the form of bubbles, followed by cooling to 100 °C by standing. To glycerol, 1.72 g (6.20 mM) of 6-(5-trifluoromethylpyridine-2-yl)-1,2,3,4-tetrahydronaphthalene and 0.50 g (1.02 mM) of Iridium (III) acetylacetonate were added, followed by heat-refluxing for 7 hours under stirring in nitrogen gas stream.

After the reaction, the reaction mixture was cooled to room temperature and poured into 100 ml of 1N-HCl. The resultant precipitate was recovered by filtration and washed with water, followed by washing with acetone and purification by silica gel column chromatography (eluent: chloroform) to obtain 0.11 g of Iridium (III) tris[6-(5-trifluoromethylpyridine-2-yl)-1,2,3,4-tetrahydronaphthalene] (orange powder) (Yield: 10.5 %).

[0170]



[0171]

As described hereinabove, the metal coordination compound (material (2)) used in the present invention provides a higher phosphorescence efficiency and a shorter phosphorescence life and allows control of its emission wavelength by appropriately modifying the alkylene group Y or substituents X1 and X2, thus being suitable as a luminescent material for EL device.

[0172]

As a result, the EL device (luminescence device) having an organic layer containing the metal coordination compound according to the present invention is an excellent device which exhibits a high efficiency luminescence, a high luminance for a long period, and a less luminescence deterioration in energized state.

[0173]

[Effect of the Invention]

As described hereinabove, the metal coordination compound used in the present invention provides a higher phosphorescence efficiency and a shorter phosphorescence life and allows control of its emission wavelength by changing the substituents, thus being suitable as a luminescent material for EL device.

As a result, the luminescence device having

an organic layer containing the metal coordination compound according to the present invention is an excellent device which exhibits a high efficiency luminescence, a high luminance for a long period, and a less luminescence deterioration in energized state.

[Brief Description of the Drawings]

Figure 1 is a schematic (sectional) view of a layer structure of a luminescence device according to the present invention.

Figure 2 is a graph showing a relationship between a Hammett's substitution constant  $\sigma$  and a peak (maximum) emission wavelength in the metal coordination compound of the present invention.

[Explanation of Symbols]

- 11: metal electrode
- 12: luminescence layer
- 13: hole-transporting layer
- 14: transparent electrode
- 15: transparent substrate
- 16: electron-transporting layer
- 17: exciton diffusion-prevention layer



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FIG. 1

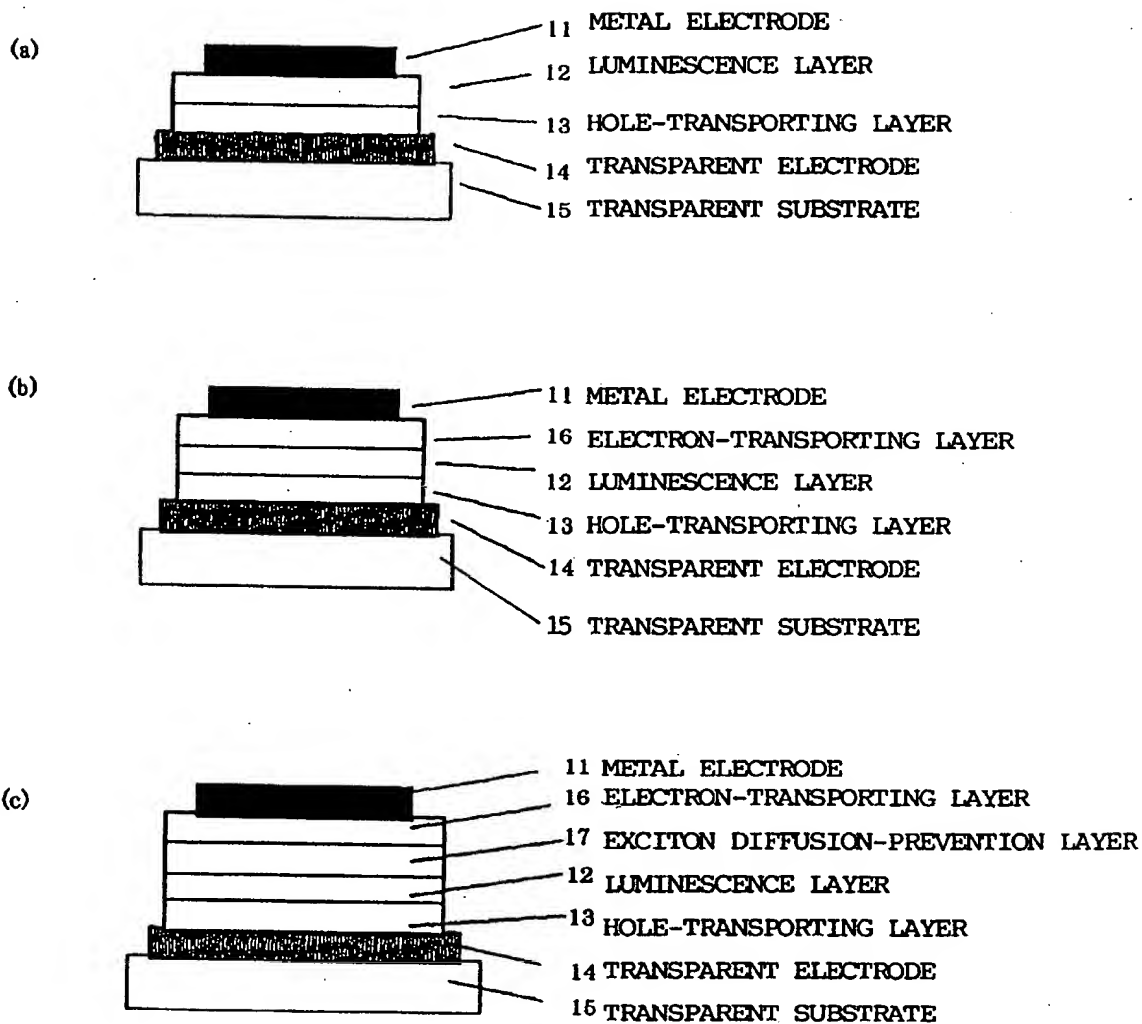


FIG. 2

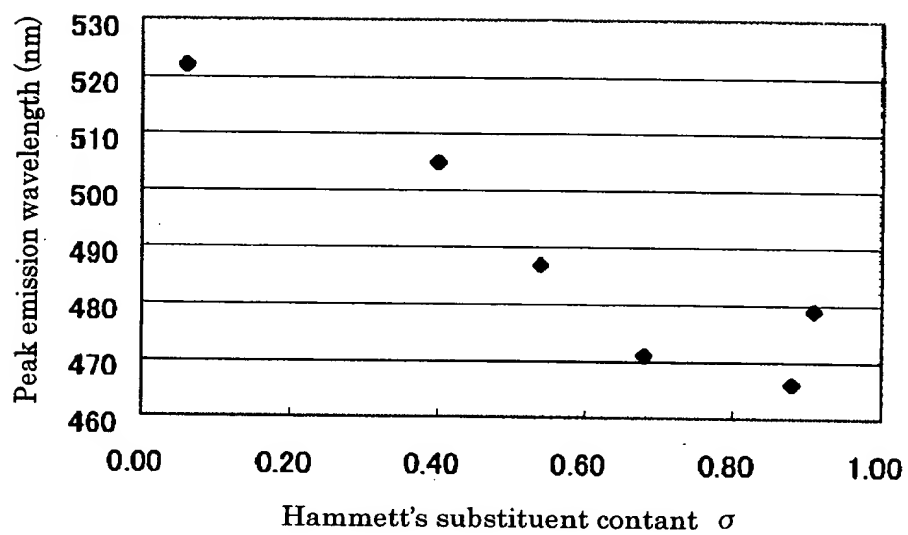
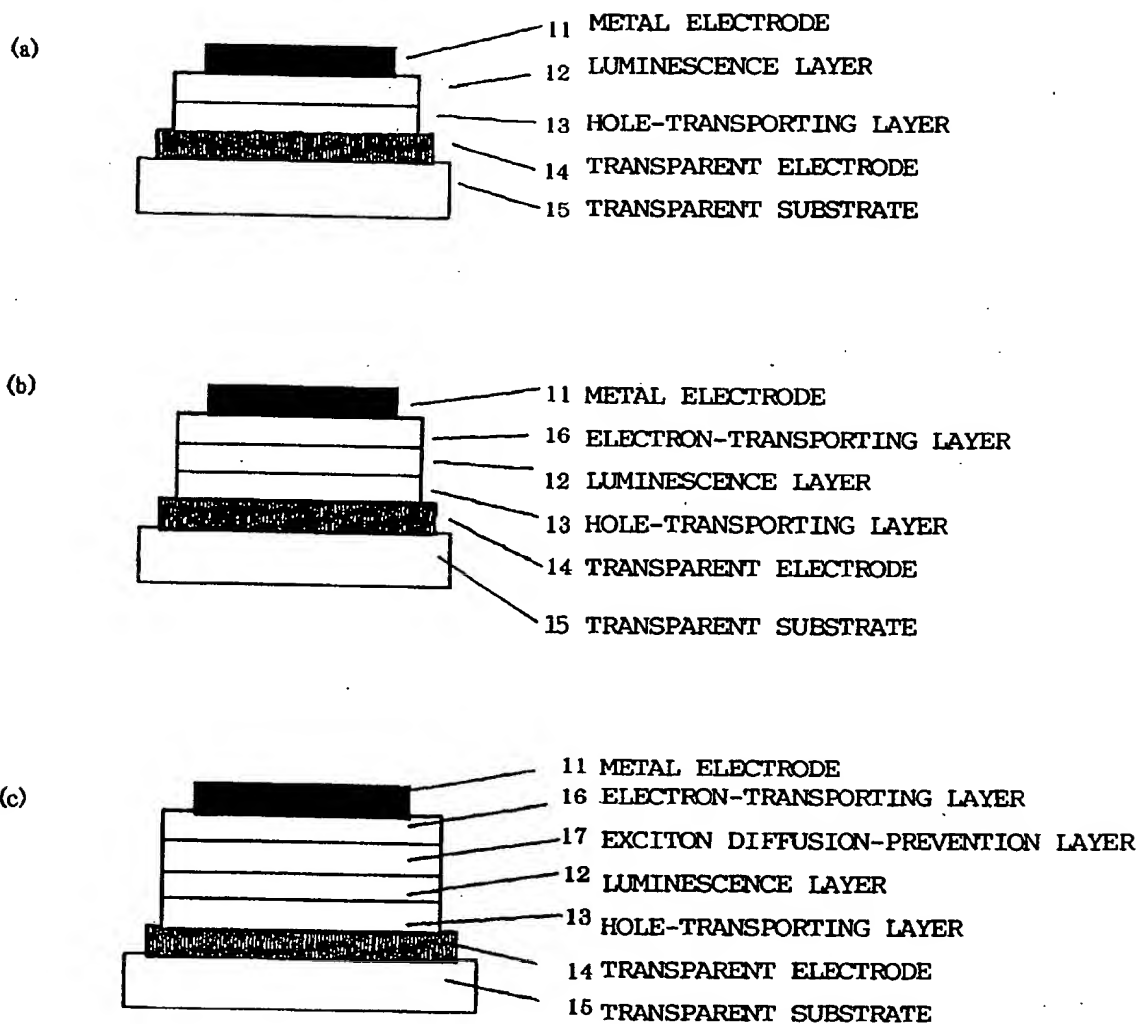


FIG. 1



[Document Name]

Abstract

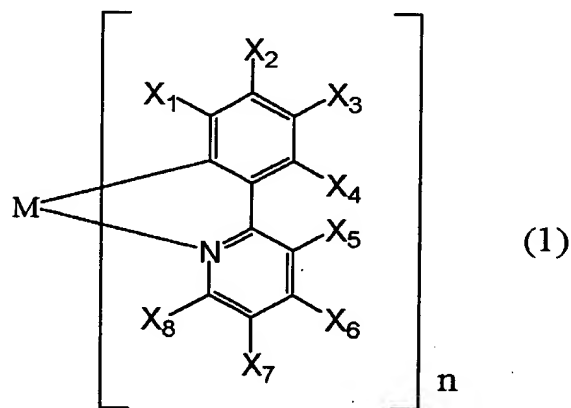
[Abstract]

[Problem]

To provide a luminescence device which produces high-efficiency luminescence, maintains a high brightness for a long period of time, and exhibits less deterioration in energized state.

[Solution Means]

A luminescence device including an organic compound layer which contains a metal coordination compound having a partial structure represented by the following formula (1):



[Selected Figure]

Figure 1